

Percent, Actual, and Theoretical Yield

Suppose a chemist reacts 100 g of CaCO_3 with 100 g of HCl to produce CaCl_2 , H_2O , and CO_2 . The chemist finds that 100 g of CaCl_2 was produced. What is the theoretical yield of CaCl_2 in grams?

$$\begin{aligned} \text{CaCO}_3 + 2\text{HCl} &\rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2 \\ 100 \text{ g CaCO}_3 &\times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{110.98 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 110.98 \text{ g CaCl}_2 \end{aligned}$$

100 g of CaCl_2 was produced. What is the theoretical yield of CaCl_2 in grams?

$$\frac{100 \text{ g CaCl}_2}{110.98 \text{ g CaCl}_2} \times 100 = 90.0\%$$

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Percent Actual And Theoretical Yield Answer Key

Morris Hein, Susan Arena, Cary Willard

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Record began publication in 1873 Debates for sessions prior to 1873 are recorded in The Debates and Proceedings in the Congress of the United States 1789 1824 the Register of Debates in Congress 1824 1837 and the Congressional Globe 1833 1873 1975 NASA Authorization United States. Congress. House. Committee on Science and Astronautics, 1974

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