

Chapter 6 Study Guide

The Periodic Table and Periodic Law

1. Explain how Mendeleev's periodic table was in error. Arranging the elements by mass resulted in several elements being placed in groups of elements with different properties. How was the error fixed? They were placed in order by increasing atomic number.
2. How was Mendeleev's periodic table organized? In order of increasing atomic mass w/ columns with similar properties.
3. What is the periodic law? Periodic repetition of chemical and physical properties of the elements when they are arranged by increasing atomic number.
4. Identify each of the following as a metal, nonmetal, or metalloid:
 - a. Oxygen NM
 - b. Barium M
 - c. Germanium M-oid
 - d. Iron M
 - e. Neon NM
 - f. Tungsten M
5. Describe the general characteristics of metals. Good conductors of heat & electricity, malleable (pounded into thin sheets), & ductile (drawn into wires).
6. Match each numbered item on the right with the lettered item that is on the left.
 1. Alkali metals b. Group 1
 2. Halogens d. Group 17
 3. Alkaline earth metals c. Group 2
 4. Noble gases a. Group 18
 - a. Group 18
 - b. Group 1
 - c. Group 2
 - d. Group 17
7. A shiny solid element also is ductile. What side of the periodic table is it likely to be found? left.
8. What are the general properties of a metalloid? Those of both metals and nonmetals. List 3 metalloids. Silicon, antimony & germanium.
9. What is the purpose of the heavy stair-step line on the periodic table? Divides metals and nonmetals.
10. Give the chemical symbol of each of the following elements:
 - a. The two elements that are liquids at room temperature mercury and bromine
 - b. The noble gas with the greatest atomic mass radon
 - c. Any metal from group 14 tin or lead
 - d. Any inner transition metal any lanthanide or actinide
11. Explain why the radius of an atom cannot be measured directly. Electron cloud based on probability and does not have a clearly defined edge.
12. Given any two elements within a group, is the element with the larger atomic number likely to have a larger or smaller atomic radius than the other element? larger.
13. Which elements are characterized as having their d orbitals filled with electrons as you move left-to-right across a period? Transition metals.
14. An element forms a negative ion when ionized. On what side of the periodic table is the element located? right.

Periodic Table Packet Answers

N Colangelo

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Periodic Table Packet Answers

1. Understanding the eBook Periodic Table Packet Answers
 - The Rise of Digital Reading Periodic Table Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Packet Answers
 - Personalized Recommendations
 - Periodic Table Packet Answers User Reviews and Ratings
 - Periodic Table Packet Answers and Bestseller Lists
5. Accessing Periodic Table Packet Answers Free and Paid eBooks

- Periodic Table Packet Answers Public Domain eBooks
 - Periodic Table Packet Answers eBook Subscription Services
 - Periodic Table Packet Answers Budget-Friendly Options
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 - Periodic Table Packet Answers Compatibility with Devices
 - Periodic Table Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Periodic Table Packet Answers
 - Highlighting and Note-Taking Periodic Table Packet Answers
 - Interactive Elements Periodic Table Packet Answers
8. Staying Engaged with Periodic Table Packet Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Packet Answers
9. Balancing eBooks and Physical Books Periodic Table Packet Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Packet Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Packet Answers
- Setting Reading Goals Periodic Table Packet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Packet Answers
- Fact-Checking eBook Content of Periodic Table Packet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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