

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu **K** Ni **Br**

- a. Explain why you made these choices: **All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.**

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu **K** Ni **Br**

- a. Explain why you made these choices: **All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.**

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu **K** Ni **Br**

- a. Explain why you made these choices: **All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.**

4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

5. a. O C **Be** **Ne** **Same Period**

- b. Na Rb **Fr** **H** **Same Group**

- c. **Pb** **C** Sn Si **Same Group**

- d. Au W S **Fr** **Ne** Zn **Challenge**

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

- a. O C **Be** **Ne** **Same Period**

- b. Na Rb **Fr** **H** **Same Group**

- c. **Pb** **C** Sn Si **Same Group**

- d. Au W S **Fr** **Ne** Zn **Challenge**

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**Robert C. Smoot, Smoot, Richard G.
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