

Worksheet 1: Periodic properties and variation of properties.

1. The following table represents the first three periods of the modern periodic table. Study the table and answer the questions that follow:

1 H 1.008							2 He 4.003
3 Li 6.941	4 Be 9.012	5 B 10.81	6 C 12.01	7 N 14.01	8 O 16.00	9 F 18.998	10 Ne 20.18
11 Na 22.99	12 Mg 24.31	13 Al 26.98	14 Si 28.09	15 P 30.97	16 S 32.06	17 Cl 35.45	18 Ar 39.95
19 K 39.10	20 Ca 40.08						

- (i) Write the formula of the sulphate of the element with atomic number 13.
- (ii) What type of bonding will be present in the oxide of the element with atomic number 1?
- (iii) Name the elements which have the highest ionization potential.
- (iv) How many electrons are present in the valence shell of the element with the atomic number 18?
- (v) What is the name given to the energy released when an atom in its isolated gaseous state accepts an electron to form an anion?
- (vi) What is the electronic configuration of the element in the third period which gains one electron to change into an anion?

2. Fill in the blanks:

- i) The atomic size _____ as we move from left to right across the period, because the _____ increases but the _____ remains the same.
- (ii) Choose the word or phrase from the brackets which correctly completes each of the following statements:
- (iii) The element below sodium in the same group would be expected to have a _____

Periodic Variation Of Properties Answer Key

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Table of Contents Periodic Variation Of Properties Answer Key

1. Understanding the eBook Periodic Variation Of Properties Answer Key
 - The Rise of Digital Reading Periodic Variation Of Properties Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Variation Of Properties Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Variation Of Properties Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Variation Of Properties Answer Key

- Personalized Recommendations
- Periodic Variation Of Properties Answer Key User Reviews and Ratings
- Periodic Variation Of Properties Answer Key and Bestseller Lists
- 5. Accessing Periodic Variation Of Properties Answer Key Free and Paid eBooks
 - Periodic Variation Of Properties Answer Key Public Domain eBooks
 - Periodic Variation Of Properties Answer Key eBook Subscription Services
 - Periodic Variation Of Properties Answer Key Budget-Friendly Options
- 6. Navigating Periodic Variation Of Properties Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Variation Of Properties Answer Key Compatibility with Devices
 - Periodic Variation Of Properties Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Variation Of Properties Answer Key
 - Highlighting and Note-Taking Periodic Variation Of Properties Answer Key
 - Interactive Elements Periodic Variation Of Properties Answer Key
- 8. Staying Engaged with Periodic Variation Of Properties Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Variation Of Properties Answer Key
- 9. Balancing eBooks and Physical Books Periodic Variation Of Properties Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Variation Of Properties Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Variation Of Properties Answer Key
 - Setting Reading Goals Periodic Variation Of Properties Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Variation Of Properties Answer Key

- Fact-Checking eBook Content of Periodic Variation Of Properties Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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