

### 3 • Molecules & Compounds

#### Mole Calculations - Difficulty Level 1

$$1 \text{ mole} = 6.02 \times 10^{23} \text{ molecules} = 22.4 \text{ L (@ STP)}$$

1. Calculate the mass of 1.58 moles  $\text{CH}_4$ . [molar mass  $\text{CH}_4$  = 16.0 g/mol]

G: 1.58 moles  $\text{CH}_4$

D: ? g  $\text{CH}_4$

$$\frac{1.58 \text{ moles } \text{CH}_4}{1} \times \frac{16.0 \text{ g } \text{CH}_4}{1 \text{ mole } \text{CH}_4} = 25.28 \approx \boxed{25.3 \text{ g } \text{CH}_4}$$

2. What volume will 7.29 moles of  $\text{CO}_2$  gas occupy at STP?

G: 7.29 moles  $\text{CO}_2$

D: ? L  $\text{CO}_2$

$$\frac{7.29 \text{ moles } \text{CO}_2}{1} \times \frac{22.4 \text{ L } \text{CO}_2}{1 \text{ mole } \text{CO}_2} = 163.376 \approx \boxed{163 \text{ L } \text{CO}_2}$$

3. How many molecules are there in a 0.00583 mole sample of  $\text{H}_2\text{O}$ ?

G: 0.00583 moles  $\text{H}_2\text{O}$

D: ? molecules  $\text{H}_2\text{O}$

$$\frac{0.00583 \text{ moles } \text{H}_2\text{O}}{1} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{H}_2\text{O}} = \boxed{3.51 \times 10^{21} \text{ molecules}}$$

4. What mass of  $\text{CO}_2$  gas occupies a volume of 100. Liters at STP? [molar mass  $\text{CO}_2$  = 44.0 g/mol]

G: 100. Liters  $\text{CO}_2$

D: ? g  $\text{CO}_2$

$$\frac{100. \text{ Liters } \text{CO}_2}{22.4 \text{ L } \text{CO}_2} \times \frac{1 \text{ mol } \text{CO}_2}{1} \times \frac{44.0 \text{ g } \text{CO}_2}{1 \text{ mol } \text{CO}_2} = \boxed{196 \text{ g } \text{CO}_2}$$

5. How many molecules are in a 35.0 gram sample of  $\text{H}_2\text{O}$ ? [molar mass  $\text{H}_2\text{O}$  = 18.0 g/mol]

G: 35.0 g  $\text{H}_2\text{O}$

D: ? molecules  $\text{H}_2\text{O}$

$$\frac{35.0 \text{ g } \text{H}_2\text{O}}{18.0 \text{ g } \text{H}_2\text{O}} \times \frac{1 \text{ mol } \text{H}_2\text{O}}{1} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol } \text{H}_2\text{O}} = \boxed{1.17 \times 10^{24} \text{ molecules}}$$

6. What volume will  $5.25 \times 10^{22}$  molecules of  $\text{CH}_4$  occupy at STP?

G:  $5.25 \times 10^{22}$  molecules  $\text{CH}_4$

D: ? L

$$\frac{5.25 \times 10^{22} \text{ molecules } \text{CH}_4}{6.02 \times 10^{23} \text{ molecules}} \times \frac{1 \text{ mol } \text{CH}_4}{1} \times \frac{22.4 \text{ L } \text{CH}_4}{1 \text{ mol } \text{CH}_4} = \boxed{1.95 \text{ L } \text{CH}_4}$$

# Skill Practice 31 Moles And Formulas Answers

**Joan I. Senyk, Darrell D. Ebbing**



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## **Skill Practice 31 Moles And Formulas Answers** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Skill Practice 31 Moles And Formulas Answers**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we shall delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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