

% Composition, Empirical Formula and Molecular Formula Worksheet

1. A sample of an unknown compound with a mass of 0.847 g has the following composition: 50.51 % fluorine and 49.49 % iron. When this compound is decomposed into its elements, what mass of each element would be recovered?

$$F = 0.4279\text{g}$$

$$Fe = 0.4192\text{g}$$

2. What is the percent composition of a carbon – oxygen compound, given that a 95.2 g sample of the compound contains 40.8 g of carbon and 54.4 g oxygen?

$$\%C = 42.9\%$$

$$\%O = 57.1\%$$

3. Find the percent composition of a compound that contains 2.63 g of carbon, 0.370 g of hydrogen and 0.580 g of oxygen in a 3.58 g sample of the compound.

$$\%C = 73.5\%$$

$$\%H = 10.3\%$$

$$\%O = 16.2\%$$

4. What is the percent composition of compound that contains 2.7369 g of chlorine, 0.4116 g of oxygen and 0.7971 g of phosphorus in a 3.9460 g sample of the compound?

$$\%Cl = 69.359\%$$

$$\%O = 10.43\%$$

$$\%P = 20.20\%$$

5. A sample of an unknown compound with a mass of 2.876 g has the following composition: 66.07 % carbon, 6.71 % hydrogen, 4.06 % nitrogen and 23.16 % oxygen. When this compound is decomposed into its elements, what mass of each element would be recovered?

$$C = 1.900\text{g}$$

$$H = 0.193\text{g}$$

$$N = 0.117\text{g}$$

$$O = 0.666\text{g}$$

6. Determine the empirical formula of a compound containing 2.644 g of gold and 0.476 g of chlorine.

$$Au:Cl = 1:1 = AuCl$$

7. Determine the empirical formula of a compound containing 0.928 g of gallium and 0.412 g of phosphorus.

$$Ga:P = 1:1 = GaP$$

8. Find the empirical formula of a compound, given that the compound is found to be 47.9 % zinc and 52.1 % chlorine by mass.

$$Zn:Cl = 1:2 = ZnCl_2$$

9. Find the empirical formula of a compound, given that a 48.5 g sample of the compound contains 1.75 g of carbon and 46.75 g of bromine.

$$C:Br = 1:4 = CBr_4$$

Percent Composition And Empirical Formulas Answer Key

Daniel C. Tofan



Percent Composition And Empirical Formulas Answer Key:

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