

# Periodic and Groups: Valence Electrons: Mastery

## Periodic and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-1, 2A-2, 3A-3, 4A-4, 5A-5, 6A-6, 7A-7, 8A-8.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns have similar properties.

Period	Group		Group										Group	
	1A	2A	3A	4A	5A	6A	7A	8A	9A	10A	11A	12A	13A	14A
1	H	He												
2	Li	Be									B	C	N	O
3	Na	Mg									Al	Si	P	S
4	K	Ca									Ga	Ge	As	Se
5	Rb	Sr									In	Sn	Sb	Te
6	Cs	Ba									Tl	Pb	Bi	Po

Thinking like the game "Monopoly", answer that an element is in period 3 and group 1.

What element is in Period 2, and Group 16A? Period 2, Group 15, Group 16A is Chlorine (Cl). The element is oxygen.

What element is in Period 5A and Period 17?

What element is in Group 2A and Period 17?

What group would you find a chlorine in?

What group Magnesium?

## Valence Electrons

Valence Electrons are the most important electrons in an atom. Each group contains has the same number of valence electrons. Valence electrons are the electrons that are involved in chemical bonding.

Hydrogen 1



1 valence electron

Lithium 3



3 valence electrons

Group 17



7 valence electrons

Hydrogen and Lithium are both in Group 1A (Hydrogen 1, Lithium 3).

Group 16, tellurium (Te) and Group 14, silicon (Si). How many valence electrons will they have?

How about Selenium?

**Group 16:** - Elements in group 16 have a full shell of electrons. Elements in group 16 and 17 are called chalcogens. Group 16 elements are called chalcogens. "H 2A, 2A, 2A" Group 16A has a full shell of valence electrons (6). 16 valence electrons (6) have, group 16 is called the chalcogens.

# Periods And Groups Valence Electrons Masses Answers

**Marco Cascella**



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