

ELECTRONEGATIVITY

For our purposes use the definitions in the chart. The most polar bond determines the polarity of a molecule (i.e. if a compound contains one non-polar, and one polar bond the molecule, as a whole, is considered to be polar)

% Ionic Character	ΔEN	polarity
0 - 5%	0 - 0.5	non-polar
10 - 50	0.5 - 1.7	polar (covalent)
50 - 100	1.7 +	ionic

A	B	C	D	E	F	G
Molecule	Lewis structure	Draw shape. Indicate bond dipoles	ΔEN of bonds	Polarity of bonds (ignore shape)	Symmetrical molecule? (i.e. all pulls cancel out)	Polarity of molecule
1. NH_3	$\begin{array}{c} \text{H} - \text{N} - \text{H} \\ \\ \text{H} \end{array}$		$3.1 - 2.1 = 1.0$	polar	No	polar
2. N_2	$:\text{N} \equiv \text{N}:$					
3. HBr	$\text{H} - \text{Br}$					
4. OCl_2						
5. SF_6						
6. SO_2						
7. $SiCl_4$						
8. CF_3Cl			$C-F:$ $C-Cl:$			
9. XeF_4 <small>Note: the EN for Xe is 2.6</small>						
10. C_2H_4			$C-C:$ $C-H:$			

Q - which binary (two element) compound would have the greatest ΔEN ?

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John S. Phillips, Cheryl Wistrom



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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998
Wistrom, 2000

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