

AP Chemistry Chapter 14 Answers – Zumdahl

14.113

- a.
 $\text{HIO}_3 < \text{HBrO}_3$; As the electronegativity of the central atom increases, acid strength increases.
- b.
 $\text{HNO}_2 < \text{HNO}_3$; As the number of oxygen atoms attached to the central nitrogen atom increases, acid strength increases.
- c.
 $\text{HOI} < \text{HOCl}$; Same reasoning as in a.
- d.
 $\text{H}_3\text{PO}_3 < \text{H}_3\text{PO}_4$; Same reasoning as in b.

14.115

- a.
 $\text{H}_2\text{O} < \text{H}_2\text{S} < \text{H}_2\text{Se}$; As the strength of the H-X bond decreases, acid strength increases.
- b.
 $\text{CH}_3\text{CO}_2\text{H} < \text{FCH}_2\text{CO}_2\text{H} < \text{F}_2\text{CHCO}_2\text{H} < \text{F}_3\text{CCO}_2\text{H}$; As the electronegativity of neighboring atoms increases, acid strength increases.
- c.
 $\text{NH}_4^+ < \text{HONH}_3^+$; Same reason as in b.
- d.
 $\text{NH}_4^+ < \text{PH}_4^+$; Same reason as in a.

14.117

In general, metal oxides form basic solutions when dissolved in water and nonmetal oxides form acidic solutions in water.

- a.
basic; $\text{CaO(s)} + \text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(aq)}$; Ca(OH)_2 is a strong base.
- b.
acidic; $\text{SO}_2\text{(g)} + \text{H}_2\text{O(l)} \rightarrow \text{H}_2\text{SO}_3\text{(aq)}$; H_2SO_3 is a weak diprotic acid.
- c.
acidic; $\text{Cl}_2\text{O(g)} + \text{H}_2\text{O(l)} \rightarrow 2\text{HOCl(aq)}$; HOCl is a weak acid.

14.119

A Lewis base is an electron pair donor, and a Lewis acid is an electron pair acceptor.

- a.
 B(OH)_3 , acid; H_2O , base
- b.
 Ag^+ , acid; NH_3 , base
- c.

Zumdahl Chemistry Answer Key

Paul B. Kelter, Steven S. Zumdahl



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 - The Rise of Digital Reading Zumdahl Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Zumdahl Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Zumdahl Chemistry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Zumdahl Chemistry Answer Key
 - Personalized Recommendations
 - Zumdahl Chemistry Answer Key User Reviews and Ratings
 - Zumdahl Chemistry Answer Key and Bestseller Lists
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 - Zumdahl Chemistry Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Zumdahl Chemistry Answer Key
 - Highlighting and Note-Taking Zumdahl Chemistry Answer Key
 - Interactive Elements Zumdahl Chemistry Answer Key
8. Staying Engaged with Zumdahl Chemistry Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Zumdahl Chemistry Answer Key
9. Balancing eBooks and Physical Books Zumdahl Chemistry Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Zumdahl Chemistry Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Zumdahl Chemistry Answer Key
- Setting Reading Goals Zumdahl Chemistry Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Zumdahl Chemistry Answer Key
- Fact-Checking eBook Content of Zumdahl Chemistry Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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