

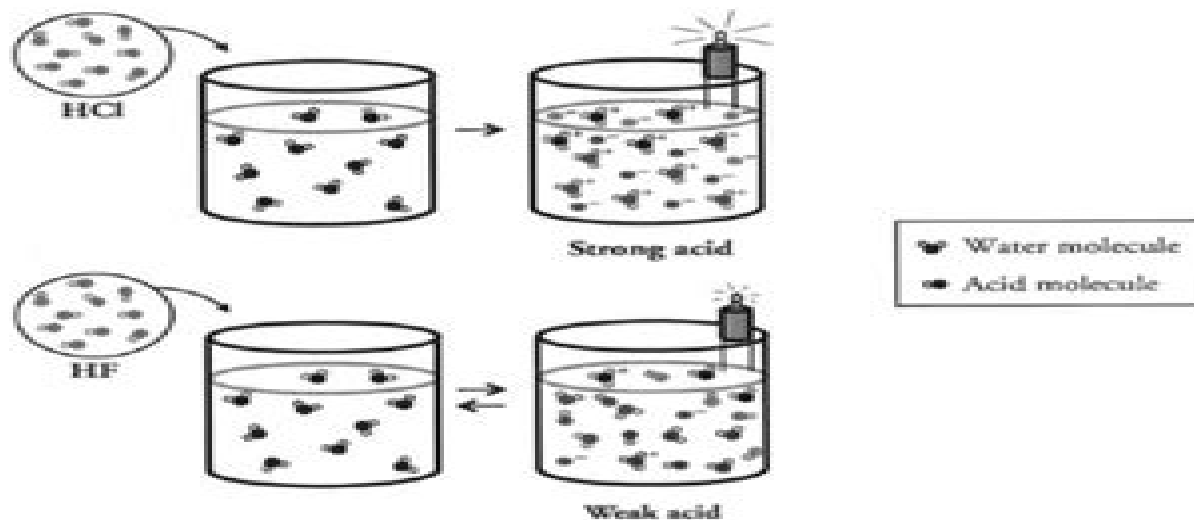
Strong versus Weak Acids

What makes a strong acid strong?

Why?

Acids are substances that surround us in our everyday life. The uses of acids range from providing essential nutrients for our bodies to dissolving metals. Some acids are safe to handle with our bare hands or even use in food preparation. Other acids will severely burn human skin. It is important to understand how these substances can all be acids and yet have such different properties.

Model 1 – Acid Strength and Conductivity



Note: Excess water molecules in the solutions are not shown.

1. Examine the symbols in Model 1. Match each symbol with its correct meaning.

- | | | | | |
|---------------|---|--------------------|---|-----------------------|
| <u>A</u> |  | <u>B</u> |  | a. Water molecule |
| <u>d</u> |  | <u>c</u> |  | b. Acid molecule |
| Hydronium ion | | conjugate base ion | | c. Conjugate base ion |
| | | | | d. Hydronium ion |

Strong Vs Weak Acids Pogil Packet Answer Key

M Mark



Strong Vs Weak Acids Pogil Packet Answer Key:

Solving Strong Acid and Base Problems, 2016 You gain a deeper understanding of acids bases and pH by working several sample problems These exercises help clarify the difference between strong and weak acids and bases and between the idea of a strong concentration versus a strong acid or base

A Test of the Theory of the Effect of Strong Acids on the Solubility of the Salts of Weak Acids David Schwartz, 1897

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