

Empirical and Molecular Formulas Worksheet

Key

Objectives:

- be able to calculate empirical and molecular formulas

Empirical Formula

- 1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\begin{array}{l} \text{C} \quad \frac{0.783\text{g C}}{12.01\text{g}} = \frac{0.065}{0.033} = 1.96 \quad 2 \\ \text{H} \quad \frac{0.196\text{g H}}{1.01\text{g}} = \frac{0.194}{0.033} = 5.88 \quad 6 \\ \text{O} \quad \frac{0.521\text{g O}}{16.00\text{g}} = \frac{0.033}{0.033} = 1 \quad 1 \end{array}$$



- 2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\text{Au} \quad \frac{89.14\text{g}}{196.97} = \frac{0.45}{0.45} = 1 \quad \text{AuO}_{1.5}$$

$$\text{O} \quad \frac{10.80\text{g}}{16\text{g}} = \frac{0.675}{0.45} = 1.5$$



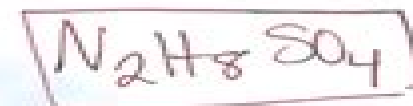
- 3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

$$\text{N} \quad \frac{21.2\text{g}}{14.01} = 1.5 \quad 2$$

$$\text{H} \quad \frac{6.1\text{g}}{1.01\text{g}} = 6.03 \quad 8$$

$$\text{S} \quad \frac{24.2\text{g}}{32} = 0.75 \quad 1$$

$$\text{O} \quad \frac{48.5\text{g}}{16.00\text{g}} = 3.03 \quad 4$$



Molecular Formula

- 4) Empirical formula of a substance is CH_2O . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \times 1 + 16) =$$

$$\frac{180}{30} = 6$$



- 5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

$$\text{C} \quad \frac{1.388}{3.585} \quad 38.7\% \rightarrow \frac{3.88\text{g}}{12.01\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{H} \quad \frac{0.345}{3.585} \quad \frac{9.62\text{g}}{1.01\text{g}} = \frac{9.52}{3.22} = 3$$

$$\text{O} \quad \frac{1.850}{3.585} \quad \frac{51.6\text{g}}{16\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{CH}_3\text{O} = \text{EFM} (12 + 3 \times 1 + 16)$$

$$\frac{62}{31.03} = 2$$



1	$\text{C}_2\text{H}_6\text{O}$
2	Au_2O_3
3	$\text{N}_2\text{H}_8\text{S}_4\text{O}_4$
4	$\text{C}_6\text{H}_{12}\text{O}_6$
5	$\text{C}_2\text{H}_6\text{O}_2$

Percent Composition And Molecular Formula Answers

IM Harris



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Chemical Calculations with Explanatory Notes, Problems, and Answers, Specially Adapted for Use in Colleges and Science Schools Richard Lloyd Whiteley, 1896 A Textbook of Organic Chemistry, 22e Arun Bahl & B S Bahl, 2019

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book is designed to meet the specific needs of underprepared students Concepts are presented only as they are needed and developed from the simple to the complex The text is divided into 18 chapters each covering some particular aspect of chemistry such as matter energy and measurement the properties of atoms description of chemical bonding study of chemical change and nuclear and organic chemistry Undergraduate students will find the book as a very valuable academic material

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Table of Contents Percent Composition And Molecular Formula Answers

1. Understanding the eBook Percent Composition And Molecular Formula Answers
 - The Rise of Digital Reading Percent Composition And Molecular Formula Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Percent Composition And Molecular Formula Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Percent Composition And Molecular Formula Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Percent Composition And Molecular Formula Answers
 - Personalized Recommendations
 - Percent Composition And Molecular Formula Answers User Reviews and Ratings
 - Percent Composition And Molecular Formula Answers and Bestseller Lists

5. Accessing Percent Composition And Molecular Formula Answers Free and Paid eBooks
 - Percent Composition And Molecular Formula Answers Public Domain eBooks
 - Percent Composition And Molecular Formula Answers eBook Subscription Services
 - Percent Composition And Molecular Formula Answers Budget-Friendly Options
6. Navigating Percent Composition And Molecular Formula Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Percent Composition And Molecular Formula Answers Compatibility with Devices
 - Percent Composition And Molecular Formula Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Percent Composition And Molecular Formula Answers
 - Highlighting and Note-Taking Percent Composition And Molecular Formula Answers
 - Interactive Elements Percent Composition And Molecular Formula Answers
8. Staying Engaged with Percent Composition And Molecular Formula Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Percent Composition And Molecular Formula Answers
9. Balancing eBooks and Physical Books Percent Composition And Molecular Formula Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Percent Composition And Molecular Formula Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Percent Composition And Molecular Formula Answers
 - Setting Reading Goals Percent Composition And Molecular Formula Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Percent Composition And Molecular Formula Answers
 - Fact-Checking eBook Content of Percent Composition And Molecular Formula Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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