

Virtual Lab – Calorimetry Experiment Worksheet

Determination of Specific Heat Capacity of Unknown Metal Using Calorimetry Experiment

1. Read the pre-lab discussion and purpose of this experiment to get a basic understanding of this virtual lab.
2. Follow the procedure given on the web page to conduct the experiment. You may want to run the demonstration first. Read over the procedure carefully before you begin.
3. Record the data that you obtained from the lab animation experiment on the table provided in page 2 and 3 and complete the calculations for specific heat capacity.
4. Determine the identity of the unknown metals by comparing your answers with those on the table in the Blackboard section of this lab.

	Ag	Al	Cu	Fe
Mass of metal	30.0 g	30.0 g	30.0 g	30.0 g
Initial temp of metal	200. °C	200. °C	200. °C	200. °C
Mass of water	80.0 g	80.0 g	80.0 g	80.0 g
Initial temp of water	20.0 °C	20.0 °C	20.0 °C	20.0 °C
Final temp of water and metal				
Specific heat of metal				

Answer the following questions:

1. What do you see happening to the temperature of the water when the hot metal block (silver, gold, iron and copper) is added to the water in the calorimeter? Is that what you would expect? Explain.
2. Which calorimeter experiment resulted in the greatest change in temperature of the water? Explain why.

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Christian G. Meyer

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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