

## Mole Conversions Worksheet

Solve the following; express your answer using the proper amount of significant figures.

1. How many moles of magnesium are in  $3.01 \times 10^{22}$  atoms of magnesium?  
**0.0500 moles of magnesium**
2. How many molecules are there in 4.00 moles of glucose,  $C_6H_{12}O_6$ ?  
 **$2.41 \times 10^{24}$  molecules of glucose**
3. How many moles are  $1.20 \times 10^{25}$  atoms of phosphorous?  
**19.9 moles of phosphorous**
4. How many moles are in 28 grams of  $CO_2$ ?  
**0.64 moles of carbon dioxide**
5. What is the mass of 5 moles of  $Fe_2O_3$ ?  
 **$8 \times 10^2$  grams of iron (III) oxide**
6. Find the number of moles of argon in 452 g of argon.  
**11.3 moles of Argon**
7. Find the mass in grams of  $2.00 \times 10^{23}$  molecules of  $F_2$ .  
**12.6 grams of difluorine**
8. Find the mass, in grams, of  $1.00 \times 10^{23}$  molecules of  $N_2$ .  
**4.65 grams of dinitrogen**
9. How many particles are there in 1.43 g of a molecular compound with a molar mass of 233 g?  **$3.70 \times 10^{21}$  particles**
10. Aspartame is an artificial sweetener that is 160 times sweeter than sucrose (table sugar) when dissolved in water. The molecular formula of aspartame is  $C_{14}H_{18}N_2O_5$ .
  - a. Calculate the molar mass of aspartame.  
**294.34 grams/mole**
  - b. How many moles are in 10 g of aspartame?  
**0.03 moles of aspartame**
  - c. How many molecules are in 5 mg of aspartame?  
 **$1 \times 10^{19}$  molecules of aspartame**
  - d. What is the mass, in grams, of 1.56 moles of aspartame?  
**459 grams of aspartame**
  - e. How many atoms of nitrogen are in 1.2 grams of aspartame?  
 **$4.9 \times 10^{21}$  atoms of nitrogen**

# Skill Practice 31 Answer Key Moles And Formulas

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