

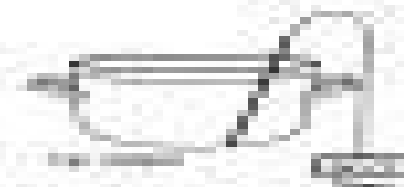
Calorimetry

What is calorimetry? How does it relate to enthalpy and temperature?

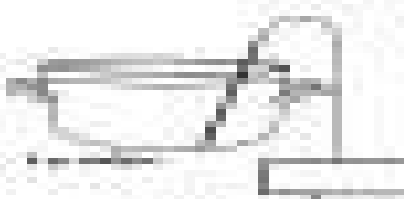
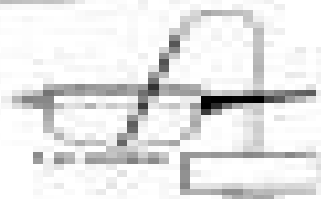
When a substance is heated, the temperature of that substance increases. Will the temperature of every other different substance be heated identical temperature increase? Will the same amount of energy be needed to raise identical temperatures for every substance at the same temperature? In this practice you will explore how mass, temperature, heat energy, and the type of substance are related.

Model 1 – A Plot of Mass

Calorimetry



How much heat energy is needed to raise 100g of water 1°C?



1. In Model 1, which substance had the most grams of water? 100g

2. Compare the graphs/curves in Model 1.

a. How many total joules of energy were added to the temperature? 100,000 J

b. How many total joules of energy were added to the temperature? 200,000 J

3. In Model 1, which substance had the highest mass? Which substance had the lowest mass? Explain your reasoning.

Why? (Explain how mass, temperature, amount of heat energy, and the type of substance are related.)

Explain how the graphs/curves in Model 1 relate to the amount of heat energy added to the temperature. (Explain how the graphs/curves in Model 1 relate to the amount of heat energy added to the temperature.)

Pogil Activities For High School Chemistry Calorimetry Packet

LL Leslie

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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POGIL Activities for High School Chemistry High School POGIL Initiative, 2012

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