

## Model 2 – Prefixes and Suffixes

Prefix	Numerical Value
mono-	1
di-	2
tri-	3
tetra-	4
penta-	5
hexa-	6
hepta-	7
octa-	8
nona-	9
deca-	10

Molecular Formula	Name of Compound
$\text{BCl}_3$	Boron trichloride
$\text{SF}_6$	Sulfur hexafluoride
$\text{IF}_7$	Iodine heptafluoride
$\text{NI}_3$	Nitrogen triiodide
$\text{N}_2\text{O}_4$	Dinitrogen tetroxide
$\text{Cl}_2\text{O}$	Dichlorine monoxide
$\text{P}_4\text{O}_{10}$	Tetraphosphorus decoxide
$\text{B}_5\text{H}_9$	Pentaboron nonahydride
$\text{Br}_3\text{O}_8$	Tribromine octoxide
$\text{ClF}$	Chlorine monofluoride

- Examine the prefixes in Model 2. Fill in the numerical value that corresponds to each prefix.
- What suffix (ending) do all the compound names in Model 2 have in common?

*They all end in "ide".*

- Carefully examine the names of the compounds in Model 2. When is a prefix NOT used in front of the name of an element?

*When the first element is one. (Examples:  $\text{IF}_7$  - iodine heptafluoride)  
 $\text{NI}_3$  - nitrogen triiodide)*

- Consider the compound NO.

- Which element, nitrogen or oxygen, would require a prefix in the molecule name? Explain your answer.

*Oxygen would require a prefix of "mono" because it is the second element and there is one oxygen. Nitrogen does not require a prefix because it is the first element and there is one nitrogen.*

- Name the molecule NO.

*Nitrogen monoxide (Notice, the "o" of mono was dropped)  
 ↳ not monoxide*

- Find two compounds in Model 2 that contain a subscript of "4" in their molecular formula.

- List the formulas and names for the two compounds.

*$\text{N}_2\text{O}_4$  - dinitrogen tetroxide ||  $\text{P}_4\text{O}_{10}$  - Tetraphosphorus decoxide*

- What is different about the spelling of the prefix meaning "four" in these two names?

*The "a" has been dropped from the end of the prefix tetra.*

# Pogil Key Naming Molecular

**George Graybill**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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