

If I have 7.7 mol of a gas at .09 atm <sup>P</sup>  
 of pressure at 56°C, <sup>T</sup> what volume does  
 this gas occupy? <sup>? V</sup>

$$56^{\circ}\text{C} + 273 = 329\text{K}$$

$$PV = nRT$$

$$\frac{(.09)(V)}{.09} = \frac{(7.7)(.0821)(329)}{.09}$$

$$\frac{PV}{P} = \frac{nRT}{P}$$

$$V = \frac{nRT}{P}$$

\*  $.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$  \* or  $8.31 \frac{\text{L} \cdot \text{kPa}}{\text{mol} \cdot \text{K}}$  or  $62.4 \frac{\text{L} \cdot \text{mmHg}}{\text{mol} \cdot \text{K}}$

# Sample Essay Questions About The Gas Laws

**Ken Addley**



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