GRAPHING TRENDS OF THE PERIODIC TABLE

	THE Y

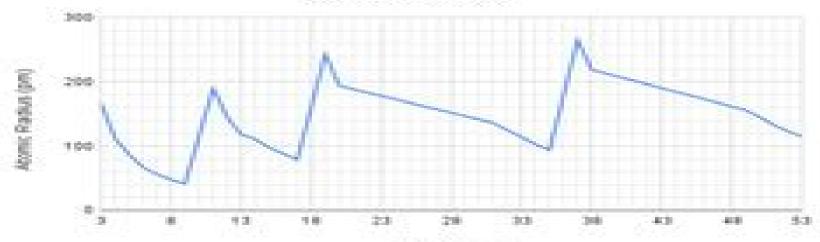
NAME		pare.
PERIOD.	Graphing Periodic Table Trends	

Directions: Using the provided graph paper, graph each of the elemental properties using a line graph. Place the property on the Y-axis and the Asomic number on the X-axis. Draw a dark, solid, vertical line on each graph where each row on the periodic table ends. (i.e. after atomic number 10). Use the graphs to analyze the trends in these properties and to answer the questions on the back.

Attornic Radius (pm)							
- 1.00 num - 3 - 1	567	Sodium - 11	190	Protessium - 19	240	Plubidium - 37	295
Blancy Blazami - 41	112	Magnessum - 52	145	California 20	104	Strombum - 28	27100
Biomom - S-	87	Alduminum - 13	1118	Gallium - 31	136	indiam : 49	156
Carbon - 6	67	Sittope - 14	211	Clermanium - 32	125	Tm - 50	7405
Nitregen - F	866	Phosphorus i 15	94	Acuerisc - 30	2004	Actionomy - 51	153
Chargen - 8	40.	Sutter - 16	888	Selection - 34	100	Tellunum - 52	1.23
Fluorine - 9	42	Chlorine - 17	20 PM 15	Binomine : 35	1946	todane - 53	1100

Romination Emergy (ILJ/Impl)							
1.8houm - 3	6.250	Soldham - 11	400	Protessoum - 19	479	Plubabum - 37	44000
Benyflum - 4	9000	Magnesture + 12	730	Calcium - 20	590	Sportfurn - 38	10000
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Atomic Radius



Adomic Number

Trends In Periodic Table Graph Answers

Wolfgang Guggemos

Trends In Periodic Table Graph Answers:

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Essential AS Chemistry for OCR Ted Lister, Janet Renshaw, 2004 Essential AS Chemistry for OCR provides clear progression with challenging material for in depth learning and understanding Written by the best selling authors of New Understanding Chemistry these texts have been written in simple easy to understand language and each double page spread is designed in a contemporary manner Fully networkable and editable Teacher Support CD ROMs are also available for this series they contain worksheets marking schemes and practical help **Johnstone Triangle** Norman Reid, 2021-01-18 Chemistry is often seen as a difficult subject to understand This book focusses on the triangle model that Alex H Johnstone developed in the early 1980s Originally conceived in the context of making chemistry more accessible to a wider range of learners the model has been applied in almost every area of education in chemistry at all stages of learning In looking at why chemistry is difficult there are two central questions Firstly does the problem relate to the nature of chemistry and secondly does it relate to the way humans gain understanding Both were found to be important and the answers to the two question were found to be connected The triangle model arose from sustained research into human learning The central finding from research is the critical role of working memory and the model rationalises so much evidence from chemistry education research as well as the repeated experiences of teachers of chemistry at all levels In order to understand chemistry it is essential to develop sound mental models of molecular reality It generates major implications for the way a chemistry curriculum should be constructed and the processes of teaching and learning in chemistry when the goal is focussed on understanding the key ideas Some of these implications are developed and pointers offered to more successful ways forward The power of the Johnstone Triangle lies in the way it offers clear directions for all involved in chemistry education It is hoped that this book will prove helpful to all involved in sharing the exciting story of the way humans have come to understand the molecular world one of the great examples of great human endeavour **Essential Laboratory Mathematics** Catherine W. Johnson, Daniel L. Timmons, Pamela E. Hall, 2009-12-03 This hands on manual with pedagogical features that draw the learner into the content offers clear and complete coverage of the mathematical topics most often used in today s clinical and medical laboratories Furthermore it provides a solid foundation for subsequent courses in the laboratory sciences

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