




## 2.2 Properties of Water

### Lesson Objectives

-  Discuss the unique properties of water.
-  Differentiate between solutions and suspensions.
-  Explain what acidic solutions and basic solutions are.

### Lesson Summary

**The Water Molecule** Water molecules ( $\text{H}_2\text{O}$ ) are polar because of an uneven distribution of electrons, creating a slight negative (–) charge in the oxygen atom and a slight positive (+) charge in each hydrogen atom. The attraction between a hydrogen atom of one water molecule and the oxygen atom of another water molecule is called a **hydrogen bond**.

- ▶ **Cohesion** is an attraction between molecules of the same substance. It causes water molecules to be drawn together, producing surface tension.
- ▶ **Adhesion** is an attraction between molecules of different substances. It causes capillary action, an effect that causes water to rise in a narrow tube against the force of gravity.

**Solutions and Suspensions** A **mixture** is a material composed of two or more elements or compounds that are physically mixed together but not chemically combined. A **solution** is a mixture in which all the components are evenly spread out: the substance dissolved is the **solute**; the substance that causes the dissolving is the **solvent**. Mixtures of water and undissolved materials are **suspensions**.

**Acids, Bases, and pH** A water molecule ( $\text{H}_2\text{O}$ ) can split apart to form a hydrogen ion ( $\text{H}^+$ ) and a hydroxide ion ( $\text{OH}^-$ ).

- ▶ The **pH scale** measures the concentration of hydrogen ions in a solution. The scale ranges from 0 to 14. Pure water has a pH of 7.
- ▶ An **acid** is any compound that forms  $\text{H}^+$  ions in solution. Acidic solutions have pH values below 7. A **base** is a compound that forms  $\text{OH}^-$  ions in solution. Basic, or alkaline, solutions have pH values above 7.
- ▶ **Buffers** are weak acids or bases that can react with strong acids or bases to prevent sudden changes in pH.

### The Water Molecule

*For Questions 1–4, write True or False on the line provided.*

- \_\_\_\_\_ 1. Water is a polar molecule.
- \_\_\_\_\_ 2. Hydrogen bonds are an example of adhesion.
- \_\_\_\_\_ 3. Covalent bonds give water a low heat capacity.
- \_\_\_\_\_ 4. A hydrogen bond is stronger than a covalent bond.

# The Properties Of Water Review And Reinforce Answers

**JA Banks**



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