

(Section-C)

[10]

Q-3 Do as directed.

(A) Circle the prime number from each of the below given list. (2)

- (1)
- 3
- , 9, 12 (2)
- 2
- , 6, 14

(B) Circle the composite number from each of the below given list. (2)

- (1) 2,
- 4
- , 11. (2)
- 21
- , 31, 41.

(C) Fill in the blanks. (4)

- (1) 1 is the smallest factor of any number.
 (2) 1 is the smallest factor of 10.
 (3) 4 is the smallest composite number.
 (4) 2 is the smallest prime number.

(D) Find the factors of given numbers by division method. (2)

$$\begin{array}{r|l}
 2 & 44 \\
 \hline
 2 & 22 \\
 \hline
 11 & 11 \\
 \hline
 & 1
 \end{array}$$

Thus, $44 = 2 \times 2 \times 11$

$$\begin{array}{r|l}
 5 & 95 \\
 \hline
 19 & 19 \\
 \hline
 & 1
 \end{array}$$

Thus $95 = 5 \times 19$

(Section-D)

[10]

Q-4 (A) State whether each of the following statement is true or false. ☐ (4)

- (1) A line segment has three points.
 (2) Infinitely many lines pass through a given point.
 (3) Line segment and ray are parts of a line.
 (4) Ray AB is denoted by \overleftrightarrow{AB}



(4)

(B) Fill in the blanks.

- (1) The fractions having the same denominator are called like fractions.
 (2) The fraction not having the same denominator are called unlike fractions.
 (3) 3 parts out of 4 equal parts of the whole is called $\frac{3}{4}$.

- (4) In $\frac{13}{16}$ the numerator is 13 and the denominator is 16.

(C) Arrange in ascending and descending order. (2)

→ Arrange in ascending order.

$$(1) \frac{6}{7}, \frac{13}{7}, \frac{30}{7}, \frac{15}{7}$$

$$\sim) \frac{6}{7}, \frac{13}{7}, \frac{15}{7}, \frac{30}{7}$$

(P.T.O.)

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 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $4\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $2\text{Al} + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + \text{Al}_2\text{SO}_4$
 $3\text{CH}_4 + 2\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$
 $8\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

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 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{GaBr}_3 + \text{NaBr} \rightarrow \text{SnO}_4 + \text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3$
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 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{NaBr} + \text{SnO}_2$
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