

Mole Conversions Practice Answer Key

Molar Mass – Calculate the Molar Mass for the following.

- What is the molar mass of Magnesium metal?

$$24.305 \text{ g/mol}$$

- Calculate the molar mass of NaCl

$$\begin{aligned} \text{Na} &: 22.990 \times 1 = 22.990 \\ \text{Cl} &: 35.453 \times 1 = 35.453 \end{aligned} \Rightarrow 58.443 \text{ g/mol}$$

- What is the molar mass of NaOH?

$$\begin{aligned} \text{Na} &: 1 \times 22.990 = 22.990 \\ \text{O} &: 1 \times 15.999 = 15.999 \\ \text{H} &: 1 \times 1.008 = 1.008 \end{aligned} \Rightarrow 39.997 \text{ g/mol}$$

Conversions – Calculate the following the conversions. Use your Mole Map to help you!

- Determine the number of atoms in 0.50 moles He.

$$0.5 \text{ mol He} \times \frac{6.02 \times 10^{23} \text{ atoms He}}{1 \text{ mol He}} = 3.01 \times 10^{23} \text{ atoms He}$$

- Determine the number of moles in 80.00 g O₂

$$0.2 \times 15.999 = 31.998 \text{ g/mol}$$

$$\Rightarrow 1 \text{ mol} = 31.998 \text{ g}$$

- Determine the volume of 0.24 mol of NO₂.

$$0.24 \text{ mol NO}_2 \times \frac{22.4 \text{ L NO}_2}{1 \text{ mol NO}_2} = 5.376 \text{ L NO}_2$$

$$80.00 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{31.998 \text{ g O}_2} = 25.00 \text{ mol O}_2$$

- Determine the number of moles in 7.5×10^{19} formula units of KNO₃

$$7.5 \times 10^{19} \text{ f.u. KNO}_3 \times \frac{1 \text{ mol KNO}_3}{6.02 \times 10^{23} \text{ f.u. KNO}_3} = 1.2458 \times 10^{-4} \text{ f.u. KNO}_3$$

- Determine the number of moles in 202.22g KNO₃

$$\begin{aligned} \text{K} &: 1 \times 39.098 = 39.098 \\ \text{N} &: 1 \times 14.007 = 14.007 \\ \text{O} &: 3 \times 15.999 = 47.997 \end{aligned} \Rightarrow 101.102 \text{ g/mol} \Rightarrow 1 \text{ mol} = 101.102 \text{ g}$$

- Determine the number of moles in 32.5L of N₂ at STP

$$32.5 \text{ L N}_2 \times \frac{1 \text{ mol N}_2}{22.4 \text{ L N}_2} = 1.451 \text{ mol N}_2$$

$$202.22 \text{ g KNO}_3 \times \frac{1 \text{ mol KNO}_3}{101.102 \text{ g KNO}_3} = 2.00 \text{ mol KNO}_3$$

- Determine the number of moles in 3.4×10^{26} formula units of NaCl.

$$3.4 \times 10^{26} \text{ f.u. NaCl} \times \frac{1 \text{ mol NaCl}}{6.02 \times 10^{23} \text{ f.u. NaCl}} = 569.784 \text{ f.u. NaCl}$$

- Determine the moles in 36.5L of carbon monoxide

$$36.5 \text{ L CO} \times \frac{1 \text{ mol CO}}{22.4 \text{ L CO}} = 1.629 \text{ mol CO}$$

The Mole Chapter Answer Key

Michael S. Leonard



The Mole Chapter Answer Key:

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Single Crystal Growth of Semiconductors from Metallic Solutions covers the four principal growth techniques currently in use for the growth of semiconductor single crystals from metallic solutions Providing an in depth review of the state of the art of each both experimentally and by numerical simulations The importance of a close interaction between the numerical and experimental aspects of the processes is also emphasized Advances in the fields of electronics and opto electronics are hampered by the limited number of substrate materials which can be readily produced by melt growth techniques such as the Czochralski and Bridgman methods This can be alleviated by the use of alternative growth techniques and in particular growth from metallic solutions The principal techniques currently in use are Liquid Phase Epitaxy Liquid Phase Electroepitaxy the Travelling Heater Method and Liquid Phase Diffusion Single Crystal Growth of Semiconductors from Metallic Solutions will serve as a valuable reference tool for researchers and graduate and senior undergraduate students in the field of crystal growth It covers most of the models developed in recent years The detailed development of basic and constitutive equations and the associated interface and boundary conditions given for each technique will be very valuable to researchers for the development of their new models Describes the fundamentals of crystal growth modelling Providing a state of the art description of the mathematical and experimental growth processes Allows reader to gain clear insight into the practical and mathematical aspects of the topic *Calculating Theoretical Yield Is Not Impossible!* Michael S. Leonard,2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill It begins with simple examples from everyday life demonstrates the importance of balancing the equation addresses the role of the mole in these computations discusses different types of liquids considers the role of significant figures and culminates with the planning of syntheses There are suggestions for further reading as well as

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Bentley's Textbook of Pharmaceutics - E-Book Sanjay Kumar Jain, Vandana Soni, 2011-08-23 This adaptation of Bentley's Textbook of Pharmaceutics follows the same goals as those of the previous edition albeit in a new look The content of the old edition has been updated and expanded and several new chapters viz Complexations Stability Testing as per ICH Guidelines Parenteral Formulations New Drug Delivery Systems and Pilot Plant Manufacturing have been included with an intention to make the book more informative for the modern pharmacists The book has six sections Section I deals with the physicochemical principles Two new chapters Complexations and ICH Guidelines for Stability Testing have been added to make it more informative Section II conveys the information regarding pharmaceutical unit operations and processes Section III describes the area of pharmaceutical practice Extensive recent updates have been included in many chapters of this section Two new chapters Parenteral Formulations and New Drug Delivery Systems have been added Section IV contains radioactivity principles and applications Section V deals with microbiology and animal products Section VI contains the formulation and packaging aspects of pharmaceuticals Pilot Plant Manufacturing concepts are added as a new chapter which may be beneficial to readers to understand the art of designing of a plant from the pilot plant model **Principles of Modern Chemistry** David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as

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