## Covalent Bonding Worksheet

Covalent bonding occurs when two or more NON METALS share electrons, attempting to attain a stable octet (8 outer electrons)in their outer shell for at least part of the time. Draw a Lewis dot diagram for each element listed. Circle the unpaired electrons that will be shared between the elements.

H<sub>2 bydrogen</sub> is distornic

H + H → H → H-H Single Bond

F<sub>2</sub> theories in distance.

3.) O2 oxygen is diatomic

O + O

\* O=O

Double Bond

4.) No nitrogen is dianomic. Is this a triple bond?

N + N

5.) BF<sub>3</sub> you need 3 fluorine atoms here

 $\mathbf{R} + \mathbf{F}$ 

6.) Ammonia NH3 hint: how many hydrogen atoms are needed?

N + H

7.) Carbon dioxide CO-

0 + c + 0

Methane CH<sub>4</sub> careful here 4 hydrogen atoms needed

C + H

9.) Dihydrogen monoxide: the most dangerous substance on the planet. It has killed more people than any other substance known to mankind!!

H + O + H

10.) SO2 hint: one pair of electrons from sulfur must be slit up for this one to work.

O + S + O

# **The Octet Rule In Covalent Bonding Worksheet Answers**

**Henry Sinclair Pickering** 

### The Octet Rule In Covalent Bonding Worksheet Answers:

**Exploration of the Structure of Atom** Chandan Sengupta, First Publication April 2019 Revised Publication October 2022 Second Revised Edition July 2023 Third Revised Edition December 2024 Published by Chandan Senguta Printed by IECIT Printing and ublication Services Resource Centre Arabinda Nagar Bankura 722101 W B India Topics Covered Atoms and Molecules Structure of Atom Bonding Mechanism and Chemical Reactions Mechanism of Bonding This book is suitable for students of Class 9 to Class 11 Students aspiring for Pre Medical Entrance Examination can also get adequate support Additional Hard Copies can also be obtained from Chandan Sukumar Sengupta Arabinda Nagar Bankura 722101 WB Write to **Holt Chemistry** Ralph Thomas Myers, 2004 Molecules and the Chemical Bond Henry A. Us for more materials Bent, 2011 MOLECULES AND THE CHEMICAL BOND Chemistry Simplified This highly original book by a famous chemistry teacher about general chemistry in a new key may change how teachers teach Atomic Theory The Mole Concept and Avogadro s Constant The Gas Laws Solving Problems in Chemical Stoichiometry The Saturation and Directional Character of Chemical Affinity The Pauli Exclusion Principle Linnett's Double Spin Set Theory Pauling's Rules of Crystal Chemistry The Octet Rule Lewis Structures for O2 NO CO SO2 and SO3 Construction of Bond Diagrams VSEPR Theory Dative Bonding Multicenter Bonding Bonding in Metals pH Calculations The Periodic Table The Energy Function and the First Law of Thermodynamics The Entropy Function and the Second Law of Thermodynamics How an Inductive Science Advances

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