

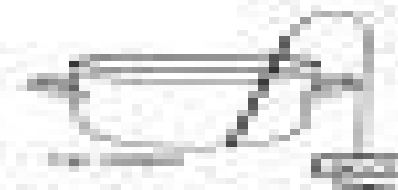
Calorimetry

What is calorimetry? How does it relate to enthalpy and temperature?

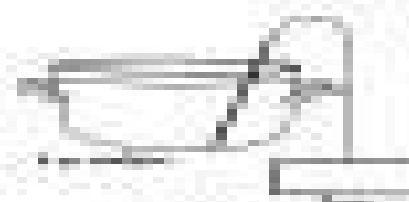
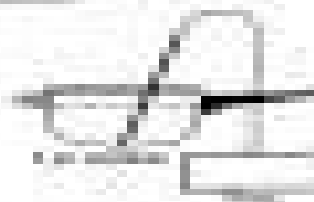
When a substance is heated, the temperature of that substance increases. Will the same amount of energy create different substances to have identical temperature increases? Will the same amount of energy be needed to create identical temperature increases in different amounts of the same substance? In this practice you will explore how mass, temperature, heat energy, and the type of substance are related.

Model 1 – A Plot of Mass

Calorimetry



How much heat energy is added to the water per minute?



1. In Model 1, which substance had the most grams of water? 200g

2. Compare the graphs/curves in Model 1.

a. How many total joules of energy were added to the temperature? 20,000 J

b. How many total joules of energy were added to the temperature? 20,000 J

3. In Model 1, which substance had the highest mass? Which substance had the lowest mass? Explain your reasoning.

Why? (Consider how much substance, amount of heat added, and the temperature change.)

Answer: The 200g of water had the highest mass and the 100g of water had the lowest mass. The 200g of water had the highest temperature change and the 100g of water had the lowest temperature change.

Pogil Activities For High School Chemistry Calorimetry Answers

Rosina Ehmann



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