

Assignment #2 – Compound Names and Formulas
Multi-valent ions only

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

1. copper (II) oxide CuO
2. mercury (I) oxide Hg₂O
3. gold (III) chloride AuCl₃
4. thallium (III) bromide TlBr₃
5. bismuth (V) oxide Bi₂O₅
6. terbium (III) oxide Tb₂O₃
7. uranium (VI) oxide UO₃
8. protactinium (IV) bromide PaBr₄
9. cerium (III) oxide Ce₂O₃
10. arsenic (V) sulphide As₂S₅

11. manganese (III) oxide Mn₂O₃
12. vanadium (II) bromide VBr₂
13. niobium(V) oxide Nb₂O₅
14. titanium (III) oxide Ti₂O₃
15. titanium (III) nitride TiN
16. iron (II) oxide FeO
17. cobalt (II) phosphide Co₃P₂
18. tin (II) oxide SnO
19. thulium (II) bromide TmBr₂
20. copper (II) bromide CuBr₂

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

1. SnCl₄ Tin (IV) Chloride
2. BiBr₅ Bismuth (V) Bromide
3. PoO₂ Polonium (IV) Oxide
4. PbI₂ Lead (II) Iodide
5. HgO Mercury (II) Oxide
6. HgCl Mercury (I) Chloride
7. Au₂O₃ Gold (III) Oxide
8. FeCl₂ Iron (II) Chloride

9. PdF₄ Palladium (IV) Fluoride
10. Os₂O₃ Osmium (III) Oxide
11. MoBr₂ Molybdenum (II) Bromide
12. VCl₅ Vanadium (V) Chloride
13. Mn₂O₃ Manganese (III) Oxide
14. CoO Cobalt (II) Oxide
15. Np₂O₃ Neptunium (III) Oxide
16. V₂O₅ Vanadium (V) Oxide

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Mike Jess



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Table of Contents Pogil Packets Naming And Formula Writing Answer Key

1. Understanding the eBook Pogil Packets Naming And Formula Writing Answer Key
 - The Rise of Digital Reading Pogil Packets Naming And Formula Writing Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Packets Naming And Formula Writing Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Packets Naming And Formula Writing Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Packets Naming And Formula Writing Answer Key
 - Personalized Recommendations
 - Pogil Packets Naming And Formula Writing Answer Key User Reviews and Ratings

- Pogil Packets Naming And Formula Writing Answer Key and Bestseller Lists
- 5. Accessing Pogil Packets Naming And Formula Writing Answer Key Free and Paid eBooks
 - Pogil Packets Naming And Formula Writing Answer Key Public Domain eBooks
 - Pogil Packets Naming And Formula Writing Answer Key eBook Subscription Services
 - Pogil Packets Naming And Formula Writing Answer Key Budget-Friendly Options
- 6. Navigating Pogil Packets Naming And Formula Writing Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Packets Naming And Formula Writing Answer Key Compatibility with Devices
 - Pogil Packets Naming And Formula Writing Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Packets Naming And Formula Writing Answer Key
 - Highlighting and Note-Taking Pogil Packets Naming And Formula Writing Answer Key
 - Interactive Elements Pogil Packets Naming And Formula Writing Answer Key
- 8. Staying Engaged with Pogil Packets Naming And Formula Writing Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Packets Naming And Formula Writing Answer Key
- 9. Balancing eBooks and Physical Books Pogil Packets Naming And Formula Writing Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Packets Naming And Formula Writing Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Packets Naming And Formula Writing Answer Key
 - Setting Reading Goals Pogil Packets Naming And Formula Writing Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Packets Naming And Formula Writing Answer Key
 - Fact-Checking eBook Content of Pogil Packets Naming And Formula Writing Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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