## Buffers

How can a solution neutralize both acids and bases?

## Why?

Buffer solutions are a mixture of substances that have a fairly constant pH regardless of addition of acid or base. They are used in medicine, industry and manufacturing to keep the pH of a solution stable regardless of other reactions that might be occurring. You have buffers in your blood, for example, because the proper function of enzymes is dependent on the pH of your blood staying in a narrow range. Eating, breathing and exercising can all alter the acidity or basicity of your blood, but the buffers in your blood help control the pH and keep all of your enzymes working properly.

## Model 1 – Comparing Solutions:

Beaker A: 100 mL Fure Water  $H_2 \bigcirc (1) + H_2 \bigcirc (1) \rightleftharpoons H_3 \bigcirc^* (aq) + \bigcirc H^-(aq)$ 

	WB WA		CA	CB
	7000 0000	-	H <sub>2</sub> O*	OH
I	8 <del></del> 8	- 89	0	-0
C			+1.0×10*	+1.0×10°
E	_		1.0×10°	1.0×10 <sup>-3</sup>

Beaker C: 100 mL of 1.00 M HNO<sub>2</sub>

$$K_a$$
 of  $HNO_2 = 4.6 \times 10^{-4}$   
 $HNO_2$  (aq)  $+ H_2O$  (l)  $\rightleftharpoons H_3O^*$  (aq)  $+ NO_2^-$  (aq)

	HNO <sub>2</sub>	-	H <sub>1</sub> O*	NO <sub>2</sub>
I	1.00		0	0
C	-30		**nc	+20
E	1.00 - x		20	×
	0.98		2.4×10 <sup>-3</sup>	2.1×10 <sup>-2</sup>

$$K_a = 4.6 \times 10^{-4} = \frac{[H_3O^*][NO_2^*]}{[HNO_2]} = \frac{x \cdot x}{1.00 - x}$$

$$4.6 \times 10^{-4} = \frac{x^2}{1.00}$$
 and  $x = 2.1 \times 10^{-3}$ 

Beaker B: 100 mL of 1.00 M HCl $HCl(ag) + H_2O(l) \rightarrow H_2O^*(ag) + Cl^*(ag)$ 

	SA	WB	CA	CB
	HCL		H <sub>3</sub> O*	CIT
1	1.00		0	0
C	-1.00		+1.00	+1.00
F	0		1.00	1.00

Beaker D: 100 mL of 1.00 M HNO<sub>2</sub> mixed with 100 mL of 1.00 M NaNO<sub>2</sub> (Be careful-dilution!!!) HNO<sub>2</sub> (aa) + H<sub>2</sub>O (1)  $\Longrightarrow$  H<sub>1</sub>O\* (aa) + NO<sub>2</sub>\* (aa)

	STATE LAND	200 E-20	The state of the s	- Name and American
	HNO <sub>2</sub>	-	H <sub>2</sub> O*	NO <sub>2</sub>
I	0.500		0	0.500
C	oc.		+50	+30
E	0.500 - x		20.	0.500 + x
	0.500		4.6×10	0.500

$$K_a = 4.6 \times 10^{-4} = \frac{[H_3O^*][NO_2^*]}{[HNO_2]} = \frac{x(0.500 + x)}{(0.500 - x)}$$
  
 $4.6 \times 10^{-4} \approx \frac{x(0.500)}{0.500} \text{ and } x = 4.6 \times 10^{-4}$ 

73.50

- For each beaker in Model 1, underseath each reaction equation, write SA (strong acid), SB (strong base), WA (weak acid), or WB (weak base) for the reactants; and write CA (conjugate acid) or CB (conjugate base) for the products.
- 2 For Beaker A and B, fill in the ICE/ICF charts to determine all equilibrium/final concentrations of the molecules or ions. ICF ("final") is used for reactions that go to completion, such as reactions with strong acids or bases. For these the change will stoichiometrically consume the limiting reactant.
- 3. For Beaker C and D, fill in the ICE charts, using "x" values as needed. Then, undermeath each ICE chart, write the K, expression and show calculations to determine the value of "x." Then, go back and modify values in the Equilibrium row so that each concentration is a numerical value. (Substitute in all "x" values.)

# **Pogil 16 Answer Key**

Cynthia J. Brame

#### **Pogil 16 Answer Key:**

**POGIL** Shawn R. Simonson, 2023-07-03 Process Oriented Guided Inquiry Learning POGIL is a pedagogy that is based on research on how people learn and has been shown to lead to better student outcomes in many contexts and in a variety of academic disciplines Beyond facilitating students mastery of a discipline it promotes vital educational outcomes such as communication skills and critical thinking Its active international community of practitioners provides accessible educational development and support for anyone developing related courses Having started as a process developed by a group of chemistry professors focused on helping their students better grasp the concepts of general chemistry The POGIL Project has grown into a dynamic organization of committed instructors who help each other transform classrooms and improve student success develop curricular materials to assist this process conduct research expanding what is known about learning and teaching and provide professional development and collegiality from elementary teachers to college professors As a pedagogy it has been shown to be effective in a variety of content areas and at different educational levels This is an introduction to the process and the community Every POGIL classroom is different and is a reflection of the uniqueness of the particular context the institution department physical space student body and instructor but follows a common structure in which students work cooperatively in self managed small groups of three or four The group work is focused on activities that are carefully designed and scaffolded to enable students to develop important concepts or to deepen and refine their understanding of those ideas or concepts for themselves based entirely on data provided in class not on prior reading of the textbook or other introduction to the topic The learning environment is structured to support the development of process skills such as teamwork effective communication information processing problem solving and critical thinking The instructor s role is to facilitate the development of student concepts and process skills not to simply deliver content to the students The first part of this book introduces the theoretical and philosophical foundations of POGIL pedagogy and summarizes the literature demonstrating its efficacy The second part of the book focusses on implementing POGIL covering the formation and effective management of student teams offering guidance on the selection and writing of POGIL activities as well as on facilitation teaching large classes and assessment The book concludes with examples of implementation in STEM and non STEM disciplines as well as guidance on how to get started Appendices provide additional resources and information about The POGIL Project **Process Oriented Guided Inquiry Learning (POGIL)** Richard Samuel Moog, 2008 POGIL is a student centered group learning pedagogy based on current learning theory This volume describes POGIL s theoretical basis its implementations in diverse environments and evaluation of student outcomes The Cambridge Handbook of Computing Education Research Sally A. Fincher, Anthony V. Robins, 2019-02-13 This is an authoritative introduction to Computing Education research written by over 50 leading researchers from academia and the industry Analytical Chemistry Iuliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical

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