

2. Refer to Model 1. For each process below indicate if the change would be endothermic or exothermic.

a. A plant making glucose from carbon dioxide and water.

*Endothermic*

b. An ice cube melting.

*Endothermic*

c. Synthesis of hydrogen peroxide from water and oxygen.

*Endothermic*

3. According to Model 1, are all exothermic reactions spontaneous? If no, provide a counter example from Model 1.

*No. The freezing of water (liquid to solid) is exothermic, but is not spontaneous above 0°C.*

4. Consider Model 1.

a. What is the symbol for the change in entropy of a system?

*S*

b. When the entropy of a system increases, what sign (+ or -) is used?

*Positive (+)*

5. According to Model 1, is there a correlation between the change in enthalpy and the change in entropy for a system? If yes, describe the correlation.

*There is no correlation between enthalpy and entropy. There are examples of exothermic reaction having increases and decreases in entropy. This also occurs for endothermic reactions.*

6. Refer to Model 1. For each process below indicate if the entropy would increase or decrease.

a. Food coloring mixing into water.

*Increase in entropy. The food coloring spreads out to occupy more space.*

b. A hot substance cooling to room temperature.

*This is a decrease in entropy. The molecular motion of the substance decreases as the substance loses energy.*

c. The respiration of glucose and oxygen to produce carbon dioxide and water.

*This is an increase in entropy. The formation of compounds from elemental oxygen increases disorder. The formation of gases from solids also increases disorder.*

7. For each of the processes in Model 1, decide as a group if the molecules in the substances have more order (were more organized) before or after the reaction. Circle the side of the reaction with more order.

8. Is there a correlation between the level of organization in the molecules and a change in entropy according to the examples in Model 1? Justify your reasoning.

*Yes, there is a correlation between level of organization and entropy. When the reactants are more organized than the products, there is an increase in entropy.  $\Delta S$  is positive.*

# Pogil High School Biology Ecological Relationships Answers

**CO Houle**



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