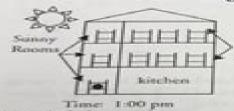
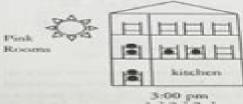
Learning Target - I understand that, when an atom is in ground state, that its electrons occupy the lowestenergy levels, sublevels, and orbitals.

Success Criteria - I can use the periodic table to determine the electron configuration for ground-state atoms.

### Model 1 - The Boarding House



Manager's Code:



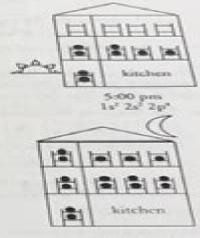
3:00 pm 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>3</sup>



Time: 7:00 pm Manager's Code: 1s 2s 2p 3s'



9:00 pm 1st 2st 2pt 3st 3pt



11:00 pm 1s' 2s' 2p' 3s' 3p'

1. Examine the boarding house diagrams in Model 1. Match each symbol below with its most likely meaning.

I. Bunk bed for boarders

II. Manager's code for the number of boarders in the house and their room assignments.

c. 1s2 2s2 2p2 3s2

III. Boarder

# **Pogil High School Chemistry Answers Electron Configurations**

N.B. Singh

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and F2 F F0 Section IV Photographs of NO and O2 molecular models Section V MRAE MO Energy Level diagrams for the diatomic structures listed above and Section VI a MRAE MO Energy Level Diagrams Template which is based on the cascade of changes in electronic distribution when an electron is forced onto a previous structure MRAE such as N2 going to N2 and NO going to NO Chapter 2 includes sections on hybridization and mechanism as it relates to polyatomic species The following sequences are given C H CH H CH2 H CH3 H CH4 saturated and tetrahedral arrangement and N H NH H NH2 H NH3 H NH4 also saturated and tetrahedral arrangement The emphasis is to show trend from atomic structure through sp sp2 and sp3 hybridization Chapter 3 is composed of selections from Atomic and Molecular Configuration 1987 Included is MRAE AUFBAU TEMPLATE of the periodic table elements expanded to two pages It is informative to compare the MRAE concept for atoms to the MRAE method for diatomic species in Chapter one In both cases the key is the changes in electronic structure when an electron is added Chemistry David John Cole-Hamilton, Iain Patterson, 2001 Chemistry Univ 1 Archie Gibb, Arthur A. Sandison, Learning and Teaching Scotland, Andrew Watson, 2000 Structures and Approximations for Electrons in Molecules David B. Cook,1978 Electronic Structure of Atoms Amin Elsersawi, 2013-02 The book presents the quantum theory of the electronic structure of atoms and focuses on the electronic structures and reactivity of atoms and molecules It shows how to draw molecules such as the oxygen and water to far more complex molecules using molecular orbital theory and hybridization of orbitals It gives quite clear picture of molecular polarity together with symmetrical and unsymmetrical distribution of an atom or molecule when developing a temporary instantaneous dipole The book provides a clear and comprehensive summary of oxidative and reductive processes Electronegativity on oxidation and reduction is also introduced Examples are provided It enables the reader to master the principles and applications of organic functional groups Readers will find information quickly and easily about alkanes alkenes alkynes and arenes Bonding with and is also introduced It explains the fundamental principles of nomenclature methods using IUPAC International Union of Pure and Applied Chemistry and enables the reader to apply it accurately and with confidence The book is replete with examples for guidance and there are extensive and complicated figures to direct the reader to nomenclature guickly It gives hands on chemistry activities with real life functions It provides clear and thorough understanding of carbohydrates polysaccharides starch and glycogen cellulose and chitin nucleotide nitrogenous hydroxyl and phosphate lipids protein ester lipoprotein glycolipid steroid mucin etc it is a useful reference for health professionals practicing physicists chemists and materials scientists **Electron Configuration**,

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