

Atomic Theory: Isotopes

Complete the following table:

Isotope Name	Atomic Number	Mass Number	Standard Atomic Notation	Number of Protons	Number of Neutrons
Carbon-14	6	14	$^{14}_6\text{C}$	6	8
Oxygen-16	8	16	$^{16}_8\text{O}$	8	8
Polonium-212	84	212	$^{212}_{84}\text{Po}$	84	128
Uranium-238	92	238	$^{238}_{92}\text{U}$	92	146
Hydrogen-2	1	2	^2_1H	1	1
Helium-4	2	4	^4_2He	2	2
Thorium-232	90	232	$^{232}_{90}\text{Th}$	90	142
Carbon-12	6	12	$^{12}_6\text{C}$	6	6
Lawrencium-257	103	257	$^{257}_{103}\text{Lr}$	103	154
Hydrogen-1	1	1	^1_1H	1	0

Questions:

- 1) How is hydrogen-1 different from hydrogen-2?

Hydrogen-1 doesn't have any neutrons. Hydrogen-2 has one neutron.

- 2) How is carbon-14 different from carbon-12?

Carbon-14 has 8 neutrons. Carbon-12 has 6 neutrons.

- 3) How do different isotopes of the same element differ?

Isotopes of the same element have different mass number because they have a different number of neutrons.

Student 1 Answers Atomic Theory

**James Johnston, Terri L. Fauber, EdD,
RT(R)(M)**



Student 1 Answers Atomic Theory:

Exploration of the Structure of Atom Chandan Sengupta, First Publication April 2019 Revised Publication October 2022 Second Revised Edition July 2023 Third Revised Edition December 2024 Published by Chandan Senguta Printed by IECIT Printing and ublication Services Resource Centre Arabinda Nagar Bankura 722101 W B India Topics Covered Atoms and Molecules Structure of Atom Bonding Mechanism and Chemical Reactions Mechanism of Bonding This book is suitable for students of Class 9 to Class 11 Students aspiring for Pre Medical Entrance Examination can also get adequate support Additional Hard Copies can also be obtained from Chandan Sukumar Sengupta Arabinda Nagar Bankura 722101 WB Write to Us for more materials

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