



Name: _____ Date: _____

Student Exploration: Colligative Properties

Background: Colligative properties are properties depending on concentration. Colligative properties include boiling point, freezing point, osmotic pressure, and vapor pressure. Colligative properties depend on solute particles.

Prior Knowledge: Questions (Do this a 1st if you're using the Gizmo.)

- What is the difference between boiling point and freezing point? Why do these differ?
- Boiling point is the temperature at which a liquid turns into a gas. Why does this happen?

- Antifreeze is added to car engines to prevent the water inside radiators from freezing. How does antifreeze work? How does antifreeze affect the boiling point?

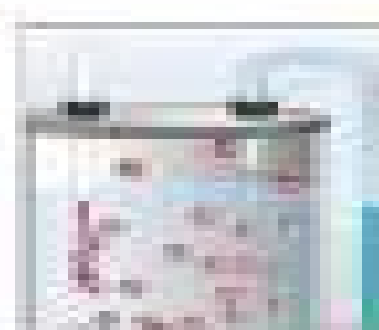
Antifreeze makes the freezing point of water go down. It also makes the boiling point go up. This is because antifreeze has a higher boiling point than water. The boiling point of water is 100°C. The boiling point of antifreeze is 110°C. This means that antifreeze will not boil at 100°C. This is why antifreeze is used in car engines. It prevents the water from freezing and the antifreeze from boiling.

Boiling Point: The temperature at which a liquid turns into a gas. Boiling point is affected by pressure. At higher pressure, the boiling point is higher. At lower pressure, the boiling point is lower. Boiling point is also affected by the amount of solute. The more solute, the higher the boiling point.

Boiling point is affected by pressure. At higher pressure, the boiling point is higher. At lower pressure, the boiling point is lower.

- What is the difference between boiling point and freezing point?

Boiling point is the temperature at which a liquid turns into a gas. Freezing point is the temperature at which a liquid turns into a solid. Boiling point is affected by pressure. Freezing point is affected by the amount of solute.



Student Exploration Colligative Properties Gizmo

Answers

R Barnett



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