

Name _____

Stoichiometry

Questions and Answers

1. What is the mole ratio in the equation $2H_2 + O_2 \rightarrow 2H_2O$?
2:1:2
2. Define stoichiometry.
Stoichiometry is the calculation of reactants and products in chemical reactions based on the balanced equation.
3. How do you determine the limiting reactant?
By calculating the amount of product formed from each reactant and identifying the reactant that forms the least product.
4. What is the theoretical yield?
The maximum amount of product that can be formed from given reactants.
5. How is percent yield calculated?
Percent Yield = (Actual Yield / Theoretical Yield) \times 100%
6. Given $N_2 + 3H_2 \rightarrow 2NH_3$, how many moles of NH_3 are produced from 3 moles of N_2 ?
6 moles
7. What are excess reactants?
Reactants that remain after the limiting reactant is completely used up.
8. Calculate the theoretical yield if the limiting reactant produces 2 moles of product.
2 moles (converted to grams if needed)
9. Define actual yield.
The amount of product actually obtained from the reaction.
10. If 5 moles of A react with 5 moles of B in the reaction $A + 2B \rightarrow AB_2$, which is the limiting reactant?
 B

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Stoichiometry

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General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

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