

Name: _____

Date: _____

Student Exploration: Calorimetry Lab

Vocabulary: calorie, calorimeter, joule, specific heat capacity

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. The Latin word *calor* means “heat,” and *meter* comes from the Greek word meaning “to measure.” What do you think a **calorimeter** does? **A tool used to measure heat.**
2. Where have you heard the word **calorie** before? What do you think a calorie is? **I've heard calories to do with nutrition and exercise.**

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the *Calorimetry Lab Gizmo™* to determine the specific heat capacities of various substances.



1. On the SIMULATION pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the GRAPH tab, and click **Play** (▶).
 - A. What was the **Final temperature** of the copper and the water? **Both 35 Celsius**
 - B. How much did the temperature of the copper change? **Decreased 55 Celsius**
 - C. How much did the temperature of the water change? **Increased 5 Celsius**
2. Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

Water has a greater heat capacity than copper because water has the least amount of change compared to copper which causes a greater resistance.

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Clifford Lane



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