

## Thermochemistry Review Part 1

**Definition Questions:** you should be able to use or apply the following terms. Be able to write complete definitions for,

energy	temperature	standard state
potential energy	heat	standard heat of formation
kinetic energy	specific heat capacity	exothermic
enthalpy	endothermic	Law of conservation of Energy
thermal	Hess's Law	isolated, closed, open system

1. The specific heat capacity of diamond is  $0.5050 \text{ J/g}^\circ\text{C}$ . How much energy is required to heat  $25.0 \text{ g}$  of diamond from  $10.5^\circ\text{C}$  to  $15.6^\circ\text{C}$ ?

$$q = m \times c \times \Delta t$$

$$= 25 \text{ g} \times 0.5050 \text{ J/g}^\circ\text{C} \times (15.6 - 10.5^\circ\text{C}) = 64.39 \text{ J}$$

2. A piece of aluminum with a mass of  $50 \text{ g}$  and an initial temperature of  $90^\circ\text{C}$  is placed into  $100 \text{ mL}$  of water at a temperature of  $25^\circ\text{C}$ . The temperature of water rises to  $31.3^\circ\text{C}$ . Determine the specific heat capacity of aluminum.

$Q \text{ lost (Al)} = Q \text{ gained (Water)}$

$$-(mc\Delta t) = +mc\Delta t$$

$$-(50 \times c \times (31.3 - 90)) = (100 \times 4.18 \times (31.3 - 25))$$

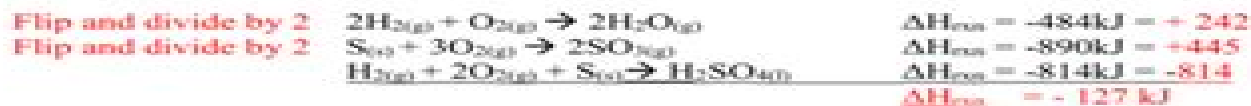
$$-(-2935 c) = 2633.4$$

$$c = 2633.4 / 2935 = 0.897 \text{ J/g}^\circ\text{C}$$

3. How much energy is absorbed by  $300 \text{ g}$  of methanol,  $\text{CH}_3\text{OH}$ , as it evaporates? Given: The molar heat of vaporization is  $35.3 \text{ kJ/mol}$ .

$q = ?$	$\Delta H = q/n$	$n \text{ CH}_3\text{OH} = m/\text{MM}$
$m = 300 \text{ g}$	$q = H \times n$	$n = 300 \text{ g} / 32 \text{ g/mol}$
$\Delta H = 35.3 \text{ kJ/mol}$	$q = (35.3 \text{ kJ} \times 9.375 \text{ mol})$	$n = 9.375 \text{ mol}$
	$q = 332.81 \text{ kJ}$	

4. What is the molar enthalpy of the formation of  $1 \text{ mol H}_2\text{SO}_{4(l)}$  given the following information?



# Thermochemistry 1 Answers

**Joseph J. Stephanos, Anthony W.  
Addison**



## Thermochemistry 1 Answers:

**Organic Chemistry Workbook** Pierre Vogel, Kendall N. Houk, 2019-11-04 Provides references and answers to every question presented in the primary Organic Chemistry textbook Successfully achieving chemical reactions in organic chemistry requires a solid background in physical chemistry Knowledge of chemical equilibria thermodynamics reaction rates reaction mechanisms and molecular orbital theory is essential for students chemists and chemical engineers The Organic Chemistry presents the tools and models required to understand organic synthesis and enables the efficient planning of chemical reactions This volume Organic Chemistry Theory Reactivity and Mechanisms in Modern Synthesis Workbook complements the primary textbook supplying the complete calculated solutions to more than 800 questions on topics such as thermochemistry pericyclic reactions organic photochemistry catalytic reactions and more This companion workbook is indispensable for those seeking clear in depth instruction on this challenging subject Written by prominent experts in the field of organic chemistry this book Works side by side with the primary Organic Chemistry textbook Includes chapter introductions and re stated questions to enhance efficiency Features clear illustrations tables and figures Strengthens reader s comprehension of key areas of knowledge Organic Chemistry Theory Reactivity and Mechanisms in Modern Synthesis Workbook is a must have resource for anyone using the primary textbook

**Hydrogen Fluoride and the Thermochemistry of Fluorine** George Thomson Armstrong, 1970 Chemistry-vol-I Dr Vijay Sarda, Dr A C Handa, Dr K K Arora, A text book on Chemistry **Structure/Reactivity and Thermochemistry of Ions** Pierre Ausloos, Sharon G.

Lias, 2012-12-06 This volume presents the proceedings of a 1986 Advanced Study Institute entitled Structure Reactivity and Thermochemistry of Ions held at Les Arcs France June 30 to July 11 1986 The format of a NATO Institute is ideally suited to in depth communications between scientists of diverse backgrounds Particularly in the field of ion physics and chemistry where on going research involves physicists physical chemists and organic chemists who use a variety of experimental and theoretical techniques it is found that in the relaxed but stimulating atmosphere of a NATO ASI each professional group provides unique insights leading to a better definition and solution of problems relating to the properties of gas phase ions This book presents chapters based on the lectures presented at the Les Arcs ASI The participants took the initiative to organize a number of specialized workshops informal discussion groups which considered questions or problem areas of particular interest The accounts of these sessions which are also included in this book make stimulating reading and include considerable useful information This Advanced Study Institute is the fourth in a series of NATO sponsored institutes devoted to the chemistry and physics of ions in the gas phase The first in 1974 in Biarritz France focussed on Interactions between Ions and Molecules

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Addison,2023-06-27 Chemical Thermodynamics and Statistical Aspects Questions to Ask in Fundamentals and Principles covers a full range of topics in macroscopic and statistical thermodynamics Every step in the book is compiled with sharp and precise attention to detail Derivations cover fundamental relationships and reinforce and extend the knowledge gained from an earlier exposure to thermodynamics The book is filled with all kinds of physics processes a variety of quantum mechanics and calculus problems involving timely mathematical functions Special emphases is given to fundamental concepts and their chemical interpretations which are essential to understanding molecular formation and reaction mechanism This book will be a useful reference source for undergraduates and postgraduates taking courses in chemistry students in chemical engineering and those in the materials sciences It will also be of value to research workers who would like an introduction to the essential principles of physical chemistry Includes detailed solutions with the necessary mathematical techniques provided for every problem Addresses problems incorporating a variety of types of chemical and physical data to illustrate the interdependence of issues Includes a Questions and Answers feature which differentiates this book from competing books in the field

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