

Thermochemistry Worksheet Key

1. How much heat, in calories, is released or absorbed if a 15.0 g sample of copper is converted from
(a) the solid to liquid state.

$$q = m \times \Delta H_{fus} = 15.0 \text{ g} \times 206 \frac{\text{J}}{\text{g}} \times \frac{1 \text{ cal}}{4.184 \text{ J}} = 738.5 = 739 \text{ cal}$$

So 739 cal is absorbed

- (b) the gas to the liquid state.

$$\begin{aligned} q &= m \times \Delta H_{con} = 15.0 \text{ g} \times -4.7 \frac{\text{kJ}}{\text{g}} \times \frac{1000 \text{ J}}{1 \text{ kJ}} \times \frac{1 \text{ cal}}{4.184 \text{ J}} = -16,849 \\ &= -1.68 \times 10^4 \text{ cal or } -16.8 \text{ kcal} \end{aligned}$$

So 16.8 kcal is released.

For copper m.p. = 1084 °C ; b.p. = 2927 °C; ΔH_{fus} = 206 J/g, ΔH_{vap} = 4.7 kJ/g

2. How much heat is absorbed or released if a 12.5 g sample of aluminum changes temperature from 25.0 to 87.5 °C? The specific heat capacity of aluminum is 0.900 J/g K

$$T_{final} = 87.5 + 273.15 = 360.65 \text{ K} \quad T_{initial} = 25.0 + 273.15 = 298.15 \text{ K}$$

$$\begin{aligned} q &= m \times c \times \Delta T = 12.5 \text{ g} \times 0.900 \frac{\text{J}}{\text{g} \cdot \text{K}} \times (360.65 - 298.15) \text{ K} \\ &= 12.5 \text{ g} \times 0.900 \frac{\text{J}}{\text{g} \cdot \text{K}} \times 62.5 \text{ K} = 703.1 = 703 \text{ J} \end{aligned}$$

3. What was the mass of a sample of water if the addition of 4.28×10^3 joules raised the temperature of the sample from 22.0 °C to 34.5 °C?

$$T_{final} = 34.5 + 273.15 = 307.65 \text{ K} \quad T_{initial} = 22.0 + 273.15 = 295.15 \text{ K}$$

T increases so endothermic so $q > 0$

$$q = m \times c \times \Delta T$$

$$\text{so } m = \frac{q}{c \times \Delta T} = \frac{4.28 \times 10^3 \text{ J}}{4.184 \frac{\text{J}}{\text{g} \cdot \text{K}} \times (307.65 - 295.15) \text{ K}} = \frac{4.28 \times 10^3 \text{ J}}{4.184 \frac{\text{J}}{\text{g} \cdot \text{K}} \times 12.5 \text{ K}} = 81.83 = 81.8 \text{ g}$$

Thermochemistry Calculations Answer

Jerzy Cioslowski



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Conceptual Chemistry Volume I For Class XI S.K. Jain & Shailesh K. Jain, 1998 Conceptual Chemistry Volume I For Class XI Calculations for A-level Chemistry E. N. Ramsden, 1995 Comprehensive mathematics foundation section Work on formulae and equations the mole volumetric analysis and other key areas is included Can be used as a course support book as well as for exam practice Best selling experienced chemistry author *Student's Guide to Fundamentals of Chemistry* Jo A. Beran, 2013-10-22 Student's Guide to Fundamentals of Chemistry Fourth Edition provides an introduction to the basic chemical principles This book deals with various approaches to chemical principles and problem solving in chemistry Organized into 25 chapters this edition begins with an overview of how to define and recognize the more common names and symbols in chemistry This text then discusses the historical development of the concept of atom as well as the historical determination of atomic weights for the elements Other chapters consider how to calculate the molecular weight of a compound from its formula This book discusses as well the characteristics of a photon in terms of its particle like properties and defines the wavelength frequency and speed of light The final chapter deals with the fundamental components of air and the classification of materials formed in natural waters This book is a valuable resource for chemistry students lecturers and instructors Thermochemistry and Its Applications to Chemical and Biochemical Systems M.A.V. Ribeiro Da Silva, 2012-12-06 Proceedings of the NATO Advanced Study Institute on Thermochemistry Today and Its Role in the Immediate Future Viano do Castelo Portugal July 5 15 1982 *Handbook of Computational Chemistry* Jerzy Leszczynski, 2012-01-14 This handbook is a guide to current methods of computational chemistry explaining their limitations and advantages and providing examples of their applications The first part outlines methods the balance of volumes present numerous important applications Fundamentals of Chemistry: A Modern Introduction (1966) Frank Brescia, 2012-12-02 Fundamentals of Chemistry A Modern Introduction focuses on the formulas processes and methodologies used in the study of chemistry The book first looks at general and historical remarks definitions of chemical terms and the classification of matter and states of aggregation The text then discusses gases Ideal gases pressure of a gas confined by a liquid Avogadro's Law and Graham's Law are described The book also discusses aggregated states of matter atoms and molecules chemical equations and arithmetic thermochemistry and chemical periodicity The text also highlights the electronic structures of atoms Quantization of electricity spectra of elements quantization of the energy of an electron associated with nucleus the Rutherford Bohr nuclear theory hydrogen atom and representation of the shapes of atomic orbitals are explained The text also highlights the types of chemical bonds hydrocarbons and their derivatives intermolecular forces solutions and chemical equilibrium The book focuses as well on ionic solutions galvanic cells and acids and bases It also discusses the structure and basicity of hydrides and oxides The reactivity of hydrides charge of dispersal and basicity effect of anionic charge inductive effect and basicity and preparation of acids are described The book is a good source of information for readers wanting to

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Quantum-Mechanical Prediction of Thermochemical Data Jerzy Cioslowski,2006-04-11 For the first time in the history of chemical sciences theoretical predictions have achieved the level of reliability that allows them to val experimental measurements in accuracy on a routine basis Only a decade ago such a statement would be valid only with severe qualifications as high level quantum chemical calculations were feasible only for molecules composed of a few atoms Improvements in both hardware performance and the level of sophistication of electronic structure methods have contributed equally to this impressive progress that has taken place only recently The contemporary chemist interested in predicting thermochemical properties such as the standard enthalpy of formation has at his disposal a wide selection of theoretical approaches differing in the range of applicability computational cost and the expected accuracy Ranging from high level treatments of electron correlation used in conjunction with extrapolative schemes to semiempirical methods these approaches have well known advantages and shortcomings that determine their usefulness in studies of particular types of chemical species The growing number of published computational schemes and their variants testing sets and performance statistics often makes it difficult for a scientist not well versed in the language of quantum theory to identify the method most adequate for his research needs

Ions in Solution J Burgess,1999-10-01 This outline of the principles and chemical interactions in inorganic solution chemistry delivers a course module in an area of considerable complexity Problems with solutions and tutorial hints to test comprehension have been added as a feature to check readers understanding and assist self study Exercises and projects are also provided to help readers deepen and extend their knowledge and understanding Inorganic solution chemistry is treated thoroughly Emphasis is placed upon NMR UV VIS IR Raman spectroscopy X ray diffraction and such topics as acid base behaviour stability constants and kinetics DDC Retrieval and Indexing Terminology Defense Documentation Center (U.S.),1975 *General Chemistry* Darrell D. Ebbing,1999 **Survival Guide to General Chemistry** Patrick E.

McMahon,Rosemary McMahon,Bohdan Khomtchouk,2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process

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Physical Chemistry for the Biosciences Raymond Chang,2005-02-11 This book is ideal for use in a one semester introductory course in physical chemistry for students of life sciences The author s aim is to emphasize the understanding of physical concepts rather than focus on precise mathematical development or on actual experimental details Subsequently only basic skills of differential and integral calculus are required for understanding the equations The end of chapter problems have both physiochemical and biological applications

General Chemistry Workbook Daniel C. Tofan,2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

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Containment Loads Scoping Calculations for Severe Nuclear Reactor Accidents ,1993

Thermochemical properties of inorganic substances I. Barin,O. Knacke,O. Kubaschewski,2013-06-29 For the practical application of thermochemistry to the development and control of tech nical processes the data for as many substances as possible are needed in conjunction with rapid and simple methods of calculating equilibrium constants heat balances and the EMF of galvanic cells For these three types of calculation the following three ther modynamic functions are suitable The Planck function the enthalpy and the Gibbs free energy which are here defined and tabulated as unambiguous functions of temperature for pure substances The first edition of the tables was published in 1973 under the title Thermochemical Properties of Inorganic Substances The present supplementary volume contains the data and functions for a further 800 inorganic substances In addition the data for about 250 substances from the

first volume have been up dated These usually small corrections produce better consistency with the data from more recent publications The comments of users and reviewers of the first volume have largely been concerned with the difference between the present thermodynamic functions and the system used in the JANAF tables the somewhat unconventional handling of heat balances adopted here the notation of cell reactions the description of non stoichiometric phases and the accuracy of the tabulated data To answer these questions and criticisms the theoretical concepts and the practical use of the tables are dealt with in more detail in the introduction following the recommendation of some reviewers

NBS Special Publication ,1968 **Publications of the National Bureau of Standards** United States. National Bureau of Standards,1983

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