

## Chapter 17 Thermochemistry Study Guide

### 17.1 – 17.2 Thermochemical Equations

1. Make the following conversions:
  - a. 444 cal to joules =  $1.86 \times 10^3 \text{ J}$
  - b. 1.8 kJ to joules =  $1.8 \times 10^3 \text{ J}$
  - c. 0.45 kJ to calories =  $1.1 \times 10^2 \text{ cal}$
2. Classify each of these processes as endothermic or exothermic:
  - a. condensing steam – **exo**
  - b. burning alcohol – **exo**
  - c. evaporating alcohol – **endo**
  - d. baking a potato – **endo**
3. The specific heat capacity for silver is  $0.24 \text{ J/g}^\circ\text{C}$ . Calculate the energy required to raise the temperature of 150.0 g Ag from 273 K to 298 K. Calculate the molar heat capacity of silver.

$$\text{Energy: } q = (150.0)(.24)(298-273) = 9.0 \times 10^2 \text{ J}$$

$$\text{Molar heat capacity: } (0.24 \text{ J/g}^\circ\text{C})(107.87 \text{ g/mol Ag}) = 26 \text{ J/mol}$$

4. It takes 585 J of energy to raise the temperature of 125.6 g Hg from  $20.0^\circ\text{C}$  to  $53.5^\circ\text{C}$ . Calculate the specific heat capacity and the molar heat capacity of Hg.

$$\text{Specific heat capacity: } C = q/m\Delta t = 585/(125.6)(53.5-20) = 0.139 \text{ J/g}^\circ\text{C}$$

$$\text{Molar heat capacity: } (0.139 \text{ J/g}^\circ\text{C})(200.59 \text{ g/mol Hg}) = 27.9 \text{ J/mol}$$

5. A 46.2-g sample of copper is heated to  $95.4^\circ\text{C}$  and then placed in a calorimeter containing 75.0 g water at  $19.6^\circ\text{C}$ . The equilibrium temperature in the calorimeter is  $21.8^\circ\text{C}$ . Calculate the specific heat capacity of copper; assuming that all the heat lost by the copper is gained by the water.

$$C = 0.203 \text{ J/g}^\circ\text{C}$$

6. A 15.0-g sample of nickel metal is heated to  $100.0^\circ\text{C}$  and dropped into 55.0 g of water, initially at  $23.0^\circ\text{C}$ . Assuming that all the heat lost by the nickel is absorbed by the water; calculate the final temperature of the nickel and the water. The specific heat of nickel is  $0.444 \text{ J/g}^\circ\text{C}$ .

$$\begin{array}{rcl} 15.0(.444)(100.0 - T_f) & = & 55.0(4.184)(T_f - 23.0) \\ 6.66T_f & = & 230.12T_f - 5292.76 \\ 5958.76 & = & 236.78 T_f \\ \mathbf{25.2^\circ\text{C}} & = & T_f \end{array}$$

# Thermochemistry Study Guide Answers

**Rachel S Tattersall**



## **Thermochemistry Study Guide Answers:**

Thermochemistry Study Guide Answers: Bestsellers in 2023 The year 2023 has witnessed a remarkable surge in literary brilliance, with numerous compelling novels enthralling the hearts of readers worldwide. Lets delve into the realm of popular books, exploring the captivating narratives that have captivated audiences this year. The Must-Read : Colleen Hoover's "It Ends with Us" This poignant tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover skillfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can prevail. Uncover the Best : Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This spellbinding historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids compelling storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Thermochemistry Study Guide Answers : Delia Owens "Where the Crawdads Sing" This captivating coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens crafts a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting. These bestselling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of engaging stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a brilliant and thrilling novel that will keep you speculating until the very end. The novel is a warning tale about the dangers of obsession and the power of evil.

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