

PRE-LABORATORY QUESTIONS
TITRATION OF A DIPROTIC ACID

Name _____

Laboratory Instructor _____ Day and time lab meets _____

1. What is a diprotic acid? (1 pt)
2. Give the balanced chemical reaction for the titration of oxalic acid with NaOH. (1 pts)



3. When titrating 25.0 mL of 0.10 M H_2SO_4 with 0.10 M NaOH, how many mL of NaOH will you have added to reach the first equivalence point? How many mL of NaOH will you have added to reach the second equivalence point? Show calculations. (4 pts)
4. A student completes a titration of an unknown diprotic acid. In their experiment, 0.79 g of the acid was dissolved in 250.0 mL of DI water. It required 13.48 mL of 1.0 M NaOH to reach the second equivalence point. What is the molar mass of the acid? Looking at the list of possible unknowns, what is likely to be the identity of the compound? Show calculations. (4pts)

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G Psacharopoulos



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