

1. What is the oxidation number of chromium in the chromate ion, CrO_4^{2-} ?
- A. +6
B. +2
C. +3
D. +8
2. What is the oxidation number assigned to manganese in KMnO_4 ?
- A. +7
B. +2
C. +3
D. +6
3. Which changes occur when Fe^{2+} is reduced?
- A. The Fe^{2+} gains electrons and its oxidation number increases.
B. The Fe^{2+} gains electrons and its oxidation number decreases.
C. The Fe^{2+} loses electrons and its oxidation number increases.
D. The Fe^{2+} loses electrons and its oxidation number decreases.
4. Given the balanced ionic equation representing a reaction:



- In this reaction, electrons are transferred from
- A. Al to Mg^{2+}
B. Al^{3+} to Mg
C. Mg to Al^{3+}
D. Mg^{2+} to Al
5. Half-reactions can be written to represent all
- A. double-replacement reactions
B. neutralization reactions
C. fusion and fission reactions
D. oxidation and reduction reactions
6. In an oxidation-reduction reaction, reduction is defined as the
- A. loss of protons
B. gain of protons
C. loss of electrons
D. gain of electrons
7. In which reaction are electrons transferred from one reactant to another reactant?
- A. $2\text{Cu}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2\text{Cu}_2\text{O}(\text{s})$
B. $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$
C. $\text{HCl}(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{NaCl}(\text{aq}) + \text{H}_2\text{O}(\text{l})$
D. $\text{H}_2\text{O}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
8. Which balanced equation represents a redox reaction?
- A. $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$
B. $\text{BaCl}_2 + \text{K}_2\text{CO}_3 \rightarrow \text{BaCO}_3 + 2\text{KCl}$
C. $\text{CuO} + \text{CO} \rightarrow \text{Cu} + \text{CO}_2$
D. $\text{HCl} + \text{KOH} \rightarrow \text{KCl} + \text{H}_2\text{O}$
9. Which metal reacts spontaneously with a solution containing zinc ions?
- A. magnesium
B. nickel
C. copper
D. silver

10. Which reaction is an example of an oxidation-reduction reaction?
- A. $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
B. $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$
C. $2\text{KOH} + \text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + 2\text{H}_2\text{O}$
D. $\text{BaCO}_3 + 2\text{HCl} \rightarrow \text{BaCl}_2 + 2\text{H}_2\text{O}$

11. Given the reaction:



Which species undergoes reduction?

- A. Al
B. Fe
C. Al^{3+}
D. Fe^{3+}
12. According to Reference Table J, which of these metals will react most readily with 1.0 M HCl to produce $\text{H}_2(\text{g})$?
- A. Cu
B. K
C. Mg
D. Zn
13. Given the balanced equation representing a redox reaction:



Which statement is true about this reaction?

- A. Each Al loses $2e^-$ and each Cu^{2+} gains $3e^-$.
B. Each Al loses $3e^-$ and each Cu^{2+} gains $2e^-$.
C. Each Al^{3+} gains $3e^-$ and each Cu loses $3e^-$.
D. Each Al^{3+} gains $3e^-$ and each Cu loses $2e^-$.
14. Given the redox reaction:
- $$\text{Cr}^{3+} + \text{Al} \rightarrow \text{Cr} + \text{Al}^{3+}$$
- As the reaction takes place, there is a transfer of
- A. electrons from Al to Cr^{3+}
B. electrons from Cr^{3+} to Al
C. protons from Al to Cr^{3+}
D. protons from Cr^{3+} to Al
15. Which half-reaction correctly represents a reduction?
- A. $\text{Ag} \rightarrow \text{Ag}^+ + e^-$
B. $\text{F}_2 \rightarrow 2\text{F}^- + 2e^-$
C. $\text{Au}^{3+} + 3e^- \rightarrow \text{Au}$
D. $\text{Fe}^{2+} + e^- \rightarrow \text{Fe}^{3+}$
16. Which equation shows conservation of both mass and charge?
- A. $\text{Cl}_2 + \text{Br}^- \rightarrow \text{Cl}^- + \text{Br}_2$
B. $\text{Cu} + 2\text{Ag}^+ \rightarrow \text{Cu}^{2+} + \text{Ag}$
C. $\text{Zn} + \text{Cr}^{3+} \rightarrow \text{Zn}^{2+} + \text{Cr}$
D. $\text{Ni} + \text{Pb}^{2+} \rightarrow \text{Ni}^{2+} + \text{Pb}$
17. Which half-reaction can occur at the anode in a voltaic cell?
- A. $\text{Ni}^{2+} + 2e^- \rightarrow \text{Ni}$
B. $\text{Sn} + 2e^- \rightarrow \text{Sn}^{2+}$
C. $\text{Zn} \rightarrow \text{Zn}^{2+} + 2e^-$
D. $\text{Fe}^{3+} \rightarrow \text{Fe}^{2+} + e^-$

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Table of Contents Regents Review Redox Worksheet

1. Understanding the eBook Regents Review Redox Worksheet
 - The Rise of Digital Reading Regents Review Redox Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Regents Review Redox Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Regents Review Redox Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Regents Review Redox Worksheet
 - Personalized Recommendations
 - Regents Review Redox Worksheet User Reviews and Ratings
 - Regents Review Redox Worksheet and Bestseller Lists
5. Accessing Regents Review Redox Worksheet Free and Paid eBooks
 - Regents Review Redox Worksheet Public Domain eBooks
 - Regents Review Redox Worksheet eBook Subscription Services
 - Regents Review Redox Worksheet Budget-Friendly Options
6. Navigating Regents Review Redox Worksheet eBook Formats

- ePub, PDF, MOBI, and More
- Regents Review Redox Worksheet Compatibility with Devices
- Regents Review Redox Worksheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Regents Review Redox Worksheet
 - Highlighting and Note-Taking Regents Review Redox Worksheet
 - Interactive Elements Regents Review Redox Worksheet
- 8. Staying Engaged with Regents Review Redox Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Regents Review Redox Worksheet
- 9. Balancing eBooks and Physical Books Regents Review Redox Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Regents Review Redox Worksheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Regents Review Redox Worksheet
 - Setting Reading Goals Regents Review Redox Worksheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Regents Review Redox Worksheet
 - Fact-Checking eBook Content of Regents Review Redox Worksheet
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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