Step 2: Balance each kind of atom other that H and O
Fe<sup>2+</sup> → Fe<sup>3+</sup> + e<sup>-</sup> Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> + 3e<sup>-</sup> → 2Cr<sup>3+</sup>

Step 3: Balance oxygen atoms by using H2O

$$Fe^{2+} \rightarrow Fe^{3+} + e^{-}$$
  $Cr_2O_7^{2-} + 3e^{-} \rightarrow 2Cr^{3+} + 7H_2O$ 

Step 4: Balance H atoms by using H+ ions

$$Fe^{2+} \rightarrow Fe^{3+} + e^{-}$$
  $14H^{+} + Cr_{2}O_{7}^{2-} + 3e^{-} \rightarrow 2Cr^{3+} + 7H_{2}O_{7}^{2-}$ 

Step 5: Use electrons as needed to obtain a charge that is balanced

Fe<sup>2+</sup> 
$$\rightarrow$$
 Fe<sup>3+</sup> + e<sup>-</sup> 14H<sup>+</sup> + Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> +3e<sup>-</sup>  $\rightarrow$  2Cr<sup>3+</sup> + 7H<sub>2</sub>O

+2 +3 -1 +14 -2 -3 +6 0

+2 +2 +9 +6

Add three electrons to the reactant side to balance the charges

$$Fe^{2+} \rightarrow Fe^{3+} + e^{-}$$
 which the the Garge  $\Theta_7^{2-} + 6e^{-} \rightarrow 2Cr^{3+} + 7H_2O$ 

# **Redox Questions And Answers**

Dr. RK Gupta

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