

Section 18.2

Reversible Reactions and Equilibrium

OBJECTIVES

- Describe how the *amounts of reactants and products change* in a chemical system at equilibrium.
- Identify three *stresses* that can change the equilibrium position of a chemical system.
- Explain what the value of K_{eq} indicates about the position of equilibrium.

Reversible Reactions

- Some reactions do not go to completion as we have assumed
 - They may be _____ – a reaction in which the conversion of reactants to products and the conversion of products to reactants occur simultaneously
 - _____ : $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$
 - _____ : $2\text{SO}_3(\text{g}) + \text{O}_2(\text{g}) \leftarrow 2\text{SO}_2(\text{g})$
 - The two equations can be combined into one, by using a _____, which tells us that it is a reversible reaction:
 $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$
 - ✓ A chemical *equilibrium* occurs, and no net change occurs in the actual amounts of the components of the system.
 - Even though the *rates* of the forward and reverse are equal, the _____ of components on both sides may not be equal
 - An equilibrium position may be shown:
- | | | | | |
|----|-----|----|-----|----|
| A | B | or | A | B |
| 1% | 99% | | 99% | 1% |
- ✓ Note the emphasis of the arrows direction
 - ✓ It depends on which side is favored; almost all reactions are reversible to some extent

Le Chatelier's Principle

- The French chemist Henri Le Chatelier (1850-1936) studied how the equilibrium position *shifts* as a result of changing conditions
 - Le Chatelier's principle: _____
-

Section 18 2 Reversible Reactions And Equilibrium

Answer Key

Sebastian Brünink



Section 18 2 Reversible Reactions And Equilibrium Answer Key:

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