

## Section 18.2

### Reversible Reactions and Equilibrium

#### OBJECTIVES

- Describe how the *amounts of reactants and products change* in a chemical system at equilibrium.
- Identify three *stresses* that can change the equilibrium position of a chemical system.
- Explain what the value of  $K_{eq}$  indicates about the position of equilibrium.

#### Reversible Reactions

- Some reactions do not go to completion as we have assumed
  - They may be \_\_\_\_\_ – a reaction in which the conversion of reactants to products and the conversion of products to reactants occur simultaneously
- \_\_\_\_\_ :  $2SO_2(g) + O_2(g) \rightarrow 2SO_3(g)$
- \_\_\_\_\_ :  $2SO_3(g) + O_2(g) \leftarrow 2SO_2(g)$
- The two equations can be combined into one, by using a \_\_\_\_\_, which tells us that it is a reversible reaction:
 
$$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$$
- ✓ A chemical *equilibrium* occurs, and no net change occurs in the actual amounts of the components of the system.
- Even though the *rates* of the forward and reverse are equal, the \_\_\_\_\_ of components on both sides may not be equal
  - An equilibrium position may be shown:
 

A	B	or	A	B
1%	99%		99%	1%
- ✓ Note the emphasis of the arrows direction
- ✓ It depends on which side is favored; almost all reactions are reversible to some extent

#### Le Chatelier's Principle

- The French chemist Henri Le Chatelier (1850-1936) studied how the equilibrium position *shifts* as a result of changing conditions
  - Le Chatelier's principle: \_\_\_\_\_
-

# Section 18 2 Reversible Reactions And Equilibrium

## Answer Key

**Sebastian Brünink**



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