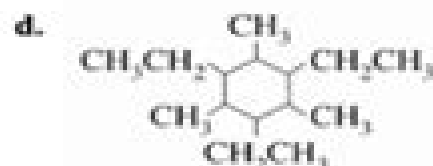


CHAPTER 21 Hydrocarbons

1. Use the IUPAC rules to name the following alkanes.



2. Draw the structure of each of the following alkanes.

a. 4-propyloctane

b. 3,4-diethylhexane

c. 2,2,4,4-tetramethylhexane

d. 1-ethyl-3-methyl-2-propylcyclopentane

3. Calculate the number of hydrogen atoms in each of the following alkanes.

a. heptane

b. cyclooctane

4. Calculate the molecular mass of a 22-carbon branched-chain alkane.

5. Chemists can analyze the composition of hydrocarbons by reacting them with copper oxide. The reaction converts carbon into carbon dioxide and hydrogen into water. Suppose 29 g of a hydrocarbon reacts to produce 88 g of CO_2 and 45 g of H_2O .

a. What are the masses of carbon and hydrogen in the hydrocarbon?

b. What is the empirical formula of the hydrocarbon?

c. If the hydrocarbon's molecular mass is 58 amu, what is its molecular formula?

6. Carbon has an electronegativity of 2.5. Hydrogen has an electronegativity of 2.2. Use these values to decide whether each of the following bonds is polar or nonpolar.

a. C-C

b. C-H

c. H-H

7. The combustion of a saturated hydrocarbon releases 657 kJ per mole of $-\text{CH}_2-$ groups and 779 kJ per mole of $-\text{CH}_3$ groups in the hydrocarbon. How much energy is released by the combustion of 1.00 L of liquid tetradecane (molecular formula $\text{C}_{14}\text{H}_{30}$), a major component of kerosene? The density of tetradecane is 0.764 g/mL.

Section 22 Hydrocarbons Answer Key

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