#### ENGAGE

Why does the atomic number but not the mass number change in positron emission?

#### SAMPLE PROBLEM 5.4 Writing a Nuclear Equation for Positron Emission

#### TRY IT FIRST

Write the balanced nuclear equation for manganese-49, which decays by emitting a positron.

#### SOLUTION GUIDE

ANALYZE THE PROBLEM	Given	Need	Connect
	Mn-49, positron	balanced nuclear	mass number, atomic
	emission	equation	number of new nucleus

STEP 1 Write the incomplete nuclear equation.

$$^{49}_{29}Mm \longrightarrow 7 + ^{9}_{47}e$$

STEP 2 Determine the missing mass number. In the equation, the mass number, 49, is equal to the sum of the mass numbers of the new nucleus and the positron.

$$49 = 7 + 0$$

$$49 - 0 = ?$$

$$49 - 0 = 49$$
 (mass number of new nucleus)

STEP 3 Determine the missing atomic number. The atomic number, 25, must equal the sum of the atomic numbers of the new nucleus and the positron.

$$25 = 7 + 1$$

$$25 - 1 = 7$$

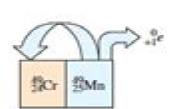
STEP 4 Determine the symbol of the new nucleus. On the periodic table, the element that has atomic number 24 is chromium, Cr. The atomic symbol for this isotope of Cr is written 50r.

SIEP 5 Complete the nuclear equation.

#### STUDY CHECK S.4.

Write the balanced nuclear equation for xenon-118, which undergoes positron emission.

#### ANSWER



# **Scientific Figures Mastering Chemistry Answers**

**Michael Seilmaier** 

**Scientific Figures Mastering Chemistry Answers:** 

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## **Table of Contents Scientific Figures Mastering Chemistry Answers**

- 1. Understanding the eBook Scientific Figures Mastering Chemistry Answers
  - The Rise of Digital Reading Scientific Figures Mastering Chemistry Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Scientific Figures Mastering Chemistry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Scientific Figures Mastering Chemistry Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Scientific Figures Mastering Chemistry Answers
  - Personalized Recommendations
  - Scientific Figures Mastering Chemistry Answers User Reviews and Ratings
  - Scientific Figures Mastering Chemistry Answers and Bestseller Lists

- 5. Accessing Scientific Figures Mastering Chemistry Answers Free and Paid eBooks
  - Scientific Figures Mastering Chemistry Answers Public Domain eBooks
  - Scientific Figures Mastering Chemistry Answers eBook Subscription Services
  - Scientific Figures Mastering Chemistry Answers Budget-Friendly Options
- 6. Navigating Scientific Figures Mastering Chemistry Answers eBook Formats
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  - Scientific Figures Mastering Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
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  - Highlighting and Note-Taking Scientific Figures Mastering Chemistry Answers
  - Interactive Elements Scientific Figures Mastering Chemistry Answers
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  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Scientific Figures Mastering Chemistry Answers
- 9. Balancing eBooks and Physical Books Scientific Figures Mastering Chemistry Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Scientific Figures Mastering Chemistry Answers
- 10. Overcoming Reading Challenges
  - o Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Scientific Figures Mastering Chemistry Answers
  - Setting Reading Goals Scientific Figures Mastering Chemistry Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Scientific Figures Mastering Chemistry Answers
  - Fact-Checking eBook Content of Scientific Figures Mastering Chemistry Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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