Review for Ch. 8 Test

8.1 Describing Chemical Reactions

 know how to write <u>word equations</u> (i.e. you have be able to name compounds, including acids, ionic, and covalent compounds from Ch. 7)

- know how to write a formula equation

 What 2 things do coefficients represent in an equation? Relative number of moles of a substance; relative number of molecules needed for a reaction to occur

- know common symbols used in a reaction

- know 4 evidences of a chemical reaction:

1. energy in the form of heat of light

2. production of a gas

3. change in color

4. formation of a precipitate

 know the <u>Law of conservation of matter</u>; matter can neither be created nor destroyed in a chemical reaction (which means, mass of reactants must equal the mass of products)

- know what a chemical equation represents: relative amounts of reactants and products

- know how to balance chemical equations (combustion reactions are particularly tricky, practice!)

Complete the equation and write a word equation for the following reaction: ZnO + FeSO₄ ⇒

ZnO + FeSO₄ ===> ZnSO₄ + FeO

Zinc oxide reacts with iron (II) sulfate to yield zinc sulfate and iron (II) oxide

2. Write a formula equation for the reaction of a solution of AgCl with potassium metal:

AgCl (aq) + K (s) - KCl (aq) + Ag (s)

3. What elements are always diatomic gases at room temperature?

All of the halogens (Group 17), H2, N2, O2

4. Can you balance these reactions:

a) 2 C₂H₆ (g) + 7 O₂ (g) - 4 CO₂ (g) + 6 H₂O (g)

b) 2 Cu₂S (s) + 3 O₂ (g) = 2 Cu₂O (s) + 2 SO₂ (g)

c) 2 Al(s) + 3 Pb(NO₃)₂ (aq) = 2 Al(NO₃)₃ (aq) + 3 Pb(s)

Section 1 1 Describing Chemical Reactions Answer Key

Sir Richard Arman Gregory (bart.), Arthur Thomas Simmons

Section 1 1 Describing Chemical Reactions Answer Key:

Numerical Solution of Time-Dependent Advection-Diffusion-Reaction Equations Willem Hundsdorfer, Jan G. Verwer, 2013-04-17 This book deals with numerical methods for solving partial differential equations PDEs coupling advection diffusion and reaction terms with a focus on time dependency A combined treatment is presented of methods for hy perbolic problems thereby emphasizing the one way wave equation methods for parabolic problems and methods for stiff and non stiff ordinary differential equations ODEs With regard to time dependency we have at tempted to present the algorithms and the discussion of their properties for the three different types of differential equations in a unified way by using semi discretizations i e the method of lines whereby the PDE is trans formed into an ODE by a suitable spatial discretization In addition for hy perbolic problems we also discuss discretizations that use information based on characteristics Due to this combination of methods this book differs substantially from more specialized textbooks that deal exclusively with nu merical methods for either PDEs or ODEs We treat integration methods suitable for both classes of problems This combined treatment offers a clear advantage On the one hand in the field of numerical ODEs highly valuable methods and results exist which are of practical use for solving time dependent PDEs something which is often not fully exploited by numerical PDE researchers Although many problems can be solved by Euler's method or the Crank Nicolson method better alter natives are often available which can significantly reduce the computational effort needed to solve practical problems Material Balances for Chemical Reacting Systems R.L. Cerro, B.G. Higgins, S. Whitaker, 2022-12-05 Written for use in the first course of a typical chemical engineering program Material Balances for Chemical Reacting Systems introduces and teaches students a rigorous approach to solving the types of macroscopic balance problems they will encounter as chemical engineers This first course is generally taken after students have completed their studies of calculus and vector analysis and these subjects are employed throughout this text Since courses on ordinary differential equations and linear algebra are often taken simultaneously with the first chemical engineering course these subjects are introduced as needed Teaches readers the fundamental concepts associated with macroscopic balance analysis of multicomponent reacting systems Offers a novel and scientifically correct approach to handling chemical reactions Includes an introductory approach to chemical kinetics Features many worked out problems beginning with those that can be solved by hand and ending with those that benefit from the use of computer software This textbook is aimed at undergraduate chemical engineering students but can be used as a reference for graduate students and professional chemical engineers as well as readers from environmental engineering and bioengineering The text features a solutions manual with detailed solutions for all problems as well as PowerPoint lecture slides available to adopting professors Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of

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