

## 12

## STOICHIOMETRY

**SECTION 12.1 THE ARITHMETIC OF EQUATIONS** (pages 353–358)

This section explains how to calculate the masses of reactants required or products formed in a stoichiometric process. It teaches you how to interpret chemical equations in terms of measuring moles, representative particles, masses, and gas volume at STP.

**► Using Everyday Equations** (pages 353–354)

1. How can you determine the quantities of reactants and products in a chemical reaction?

You must enter **2** or **balancing** **1** equals **1**

2. Quantity usually means the \_\_\_\_\_ **amount** \_\_\_\_\_ of a substance expressed in grams or moles.

3. A hot-dogger is no the fault flows 3 wheels (3H), 2 axle boards (2A), 1 top (T), 1 base (B), and 4 legs (L). Write a “balanced equation” for the construction of this hot-dogger.

**$3\text{H} + 2\text{A} + 1\text{T} + 1\text{B} + 4\text{L} = 1\text{hot-dogger}$**

**► Using Balanced Chemical Equations** (page 354)

4. In the following sentence (or on your Student Laboratory) is the calculation of quantities in chemical reactions. \_\_\_\_\_ **True**

5. Calculations using balanced equations are called \_\_\_\_\_ **stoichiometric calculations**

**► Interpreting Chemical Equations** (pages 354–357)

6. Does what element is ammonia produced? How is it used?

**Ammonia  $\text{NH}_3$  is a compound of nitrogen and hydrogen. It is used as a fertilizer.**

7. Circle the letter of the term that tells what kind of information you **CANNOT** get from a chemical equation.

- |                             |                        |
|-----------------------------|------------------------|
| a. moles                    | d. volume              |
| b. mass                     | e. number of particles |
| <b>c. size of particles</b> |                        |

# Section 12 The Arithmetic Of Equations Practice Problems

**Samuel H. WINTER**



## **Section 12 The Arithmetic Of Equations Practice Problems:**

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