

Chapter 2: Water

2.1 Water Weak Interactions in Aqueous Systems

Water's Qualities:

-Unusual properties of water caused by Hydrogen Bonds:

- 1) High Surface tension – beads up
- 2) High heat capacity
- 3) High Heat of Vaporization – a lot of energy is needed for water to boil
- 4) Expands as it freezes – the solid forms of water (ice) floats instead of sinks like most

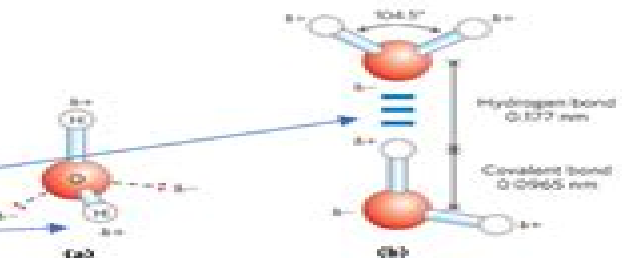
Inter/Intramolecular Forces 1-HYDROGEN BONDING

-Hydrogen bonding occurs when the hydrogen atoms of one molecule is attracted to another molecule that has Nitrogen, Oxygen, and Sulfur (NOS).

- Oxygen is the most electronegative followed by sulfur and Nitrogen.

-Picture to the right: Water Molecules

- O is more electronegative than H
- This creates a POLAR covalent bond between H and O
- The H-Bond occurs between water molecules
- O has a partial negative (-) charge
- H has a partial positive (+) charge



-Polar function groups form H-Bonds with one another and water

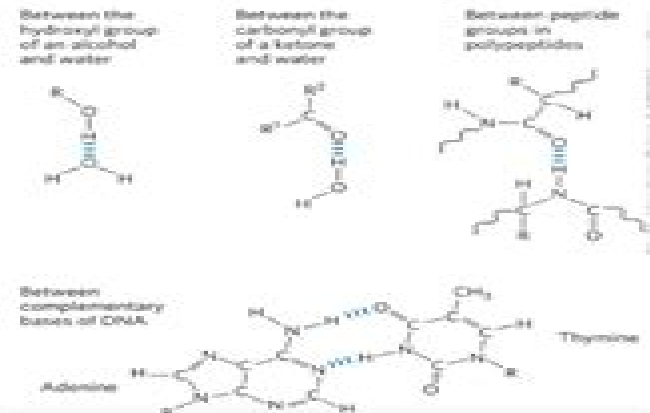


-Angle and distance affect the strength of H-Bonds



-we will see H-Bonds occur between:

- Hydroxyl group of alcohol and water
- Carbonyl group of a ketone and water
- between peptide groups and peptides
- complementary bases of DNA



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LP Steffe



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Table of Contents Section 2 Chem Study Guide Answers

1. Understanding the eBook Section 2 Chem Study Guide Answers
 - The Rise of Digital Reading Section 2 Chem Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 2 Chem Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 2 Chem Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 2 Chem Study Guide Answers
 - Personalized Recommendations
 - Section 2 Chem Study Guide Answers User Reviews and Ratings
 - Section 2 Chem Study Guide Answers and Bestseller Lists
5. Accessing Section 2 Chem Study Guide Answers Free and Paid eBooks
 - Section 2 Chem Study Guide Answers Public Domain eBooks
 - Section 2 Chem Study Guide Answers eBook Subscription Services
 - Section 2 Chem Study Guide Answers Budget-Friendly Options
6. Navigating Section 2 Chem Study Guide Answers eBook Formats

- ePub, PDF, MOBI, and More
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- Section 2 Chem Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 2 Chem Study Guide Answers
 - Highlighting and Note-Taking Section 2 Chem Study Guide Answers
 - Interactive Elements Section 2 Chem Study Guide Answers
- 8. Staying Engaged with Section 2 Chem Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 2 Chem Study Guide Answers
- 9. Balancing eBooks and Physical Books Section 2 Chem Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 2 Chem Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section 2 Chem Study Guide Answers
 - Setting Reading Goals Section 2 Chem Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 2 Chem Study Guide Answers
 - Fact-Checking eBook Content of Section 2 Chem Study Guide Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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