

Name _____

KEY

Period _____

Date _____

REVIEW QUESTIONS FOR FINAL EXAM (2nd Semester)

SECTION 12.1

Heat Transfer: $Q = mc\Delta T$ or $Q = mC(T_f - T_i)$

Specific Heat of Common Substances			
Material	Specific heat $J/kg \cdot ^\circ C$	Material	Specific heat $J/kg \cdot ^\circ C$
aluminum	900	lead	130
brass	376	nickel	240
carbon	712	silver	235
copper	385	steel	450
glass	840	water	4180
ice	2090	iron	450

1) Two solid objects are in contact as shown below.



Heat will always flow from object A to object B if object B has a lower

- a) mass b) total energy c) specific heat
d) temperature e) none of the answers

For questions 2-3: The diagram below shows two solids with their temperatures ($A=50^\circ C$, $B=80^\circ C$).

2) What occurs when the two solids are placed in contact with each other as shown?

- a) Heat energy flows from A to B. Solid A cools down.
 b) Heat energy flows from A to B. Solid A gets hotter.
c) Heat energy flows from B to A. Solid B cools down.
 d) Heat energy flows from B to A. Solid B gets hotter.

3) The net flow of heat energy will stop from one object to the other when they both have the same

- a) energy **b) temperature**
 c) specific heat d) all the answers are correct

4) How much heat is transferred when a 0.2 kg piece of copper's temperature is raised from $50^\circ C$ to $80^\circ C$?

$$Q = mc\Delta T$$

$$Q = (0.2 \text{ kg})(385 \text{ J/kg} \cdot ^\circ C)(80^\circ C - 50^\circ C) = \boxed{2310 \text{ J}}$$

5) A 0.5 kg block of lead cools down from $30^\circ C$ to $20^\circ C$. How much heat was transferred?

$$Q = (0.5 \text{ kg})(130 \text{ J/kg} \cdot ^\circ C)(20^\circ C - 30^\circ C)$$

$$\boxed{Q = -650 \text{ J}}$$

6) The heat transfer of a substance is 9400 J. The mass of this substance is 2 kg and its temperature rises from $80^\circ C$ to $100^\circ C$. Based on the specific heat that you will calculate for this substance, what is it?

$$C = \frac{Q}{m\Delta T} = \frac{9400 \text{ J}}{(2 \text{ kg})(100^\circ C - 80^\circ C)} = 235 \text{ J/kg} \cdot ^\circ C$$

SILVER

7) From the choices below, which material would be better for a pan handle?

- a) brass **376** b) silver **235** c) copper **385**
d) iron 450 e) zinc **388** f) lead **130**

8) Provide a brief description of absolute zero and their temperatures in Kelvin (K) and degrees Celsius ($^\circ C$).

The lowest temperature when all particle motion stops. $T_C = -273^\circ C$, $T_K = 0 \text{ K}$.

9) The boiling point of water is $100^\circ C$. The boiling point of water in Kelvin (K) will be

$$T_K = 273 + T_C = 273 + 100 = \boxed{373 \text{ K}}$$

10) A 1-liter of water taken out of a 5-gallon bucket full of water at $20^\circ C$ will

a) register a lower temperature than that of the water left in the bucket

b) have less thermal energy than that of the water left in the bucket

c) have less specific heat than that of the water left in the bucket

11) 243 K in degrees Celsius ($^\circ C$) is

$$T_C = T_K - 273 = 243 \text{ K} - 273 = \boxed{-30^\circ C}$$

12) The kinetic molecular theory can explain the existence of thermal energy in a substance.

- a) true** b) false

13) When using a thermometer for measuring the temperature of a substance, one should record the temperature

a) about 2 minutes after the thermometer has been placed in the substance. That is usually the time needed for the liquid inside to expand or shrink completely.

b) until thermal equilibrium has been reached

c) until all the heat from the substance has been transferred to the thermometer

d) until the thermal energy from the substance is equal to the thermal energy of the thermometer

e) all of the answers

14) A calorimeter contains 1 kg of water at $20^\circ C$. A 0.20 kg block of zinc at $120^\circ C$ is placed in the water. What is the approximate final temperature of the system?

$$T_F = \frac{m_A c_A T_{Ai} + m_B c_B T_{Bi}}{m_A c_A + m_B c_B}$$

$$T_F = \frac{(1)(4180)(20^\circ C) + (0.2)(388)(120^\circ C)}{(1)(4180) + (0.2)(388)}$$

$$T_F = \frac{83600 + 9312}{4257.6} = \boxed{21.8^\circ C}$$

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Anna N. Bukiya,Alex M. Dopico



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Cholesterol Anna N. Bukiya, Alex M. Dopico, 2022-04-26 With Cholesterol Drs Anna Bukiya and Alex Dopico have compiled a comprehensive resource on biological and clinical aspects of cholesterol spanning biophysics and biochemistry as well as the latest pharmacological discoveries employed to tackle disorders associated with abnormal cholesterol levels Early chapters on basic biology offer guidance in cholesterol lab chemistry cholesterol metabolism and synthesis molecular evolution of cholesterol and sterols cholesterol peptides and cholesterol modulation Chapters on cellular and organismal development discuss cholesterol transport in blood lipoproteins and cholesterol metabolism cholesterol detection in the blood cellular cholesterol levels hypercholesterolemia and the role of cholesterol in early human development Pathophysiological specialists consider familial hypobetalipoproteinemia critical illness and cholesterol levels coronary artery disease CESD cholesterol and viral pathology cholesterol and neurodegenerative disorders and cholesterol and substance use disorders A final section examines pharmacology of drug delivery systems targeting cholesterol related disorders cholesterol receptors cholesterol reduction statins citrate lyase cyclodextrins and clinical management Cholesterol From Biophysics and Biochemistry to Pathology and Pharmacology empowers researchers students and clinicians across various disciplines to advance new cholesterol based studies improve clinical management and drive drug discovery Ties basic biology to clinical application and drug discovery Provides methods and protocols for lab based cholesterol research and clinical testing Examines the latest pharmacological discoveries employed to tackle cholesterol related disorders Includes chapter contributions from a wide range of specialists uniting various disciplines *Ozone Hole* S.A. Abbasi, Tasneem Abbasi, 2017-03-10 This book underscores the re emergence of the ozone hole problem and deals with it in its current context of exacerbating global warming It traces the history of the ozone hole from the stage of formation of the stratospheric ozone layer millions of years ago into the late 20th century when the anthropogenic destruction of that ozone was discovered The chapters are written to bring the the reader up to the present day Factors that influence stratospheric ozone are discussed and the ways to halt ozone depletion are cataloged And more complex interrelationships are being discovered between ozone depletion and two other global concerns climate change and ocean acidification This book sheds light on the intricacy of the situation and its portants The book will be useful to students and researchers looking for a current overview of the ozone hole problem div

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