

8.2

THE NATURE OF COVALENT BONDING

KEY

Section Review

Objectives

- State a rule that usually tells how many electrons are shared to form a covalent bond
- Describe how electron dot formulas are used
- Predict when two atoms are likely to be joined by a double or a triple covalent bond
- Distinguish between a single covalent bond and other covalent bonds
- Describe how the strength of a covalent bond is related to its bond dissociation energy
- Describe how resonance structures explain bonding
- Identify some exceptions to the octet rule

Vocabulary

- | | | |
|------------------------|----------------------------|----------------------------|
| • single covalent bond | • double covalent bonds | • polyatomic ion |
| • structural formulas | • triple covalent bonds | • bond dissociation energy |
| • unshared pairs | • coordinate covalent bond | • resonance structures |

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

When atoms share electrons to gain the 1 configuration of a noble gas, the bonds formed are 2. A 3 pair of valence electrons constitutes a 4 covalent bond. Pairs of valence electrons that are not shared between atoms are called 5. Sometimes two or three pairs of electrons may be shared to give 6 covalent bonds. In some cases, only one of the atoms in a bond provides the pair of bonding electrons; this is a 7. 8 is required to break covalent bonds between atoms. The total energy required to break the bond between two covalently bonded atoms is known as the 9.

When it is possible to write two or more valid electron dot formulas for a molecule or ion, each formula is referred to as a 10.

1. stable
2. covalent
3. shared
4. single
5. unshared pairs
6. double/triple
7. coordinate covalent bond
8. energy
9. bond dissociation energy
10. resonance structure

Section 8 2 Answers The Nature Of Covalent Bonding

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Table of Contents Section 8 2 Answers The Nature Of Covalent Bonding

1. Understanding the eBook Section 8 2 Answers The Nature Of Covalent Bonding
 - The Rise of Digital Reading Section 8 2 Answers The Nature Of Covalent Bonding
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 8 2 Answers The Nature Of Covalent Bonding
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 8 2 Answers The Nature Of Covalent Bonding
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 8 2 Answers The Nature Of Covalent Bonding
 - Personalized Recommendations
 - Section 8 2 Answers The Nature Of Covalent Bonding User Reviews and Ratings
 - Section 8 2 Answers The Nature Of Covalent Bonding and Bestseller Lists

5. Accessing Section 8 2 Answers The Nature Of Covalent Bonding Free and Paid eBooks
 - Section 8 2 Answers The Nature Of Covalent Bonding Public Domain eBooks
 - Section 8 2 Answers The Nature Of Covalent Bonding eBook Subscription Services
 - Section 8 2 Answers The Nature Of Covalent Bonding Budget-Friendly Options
6. Navigating Section 8 2 Answers The Nature Of Covalent Bonding eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 8 2 Answers The Nature Of Covalent Bonding Compatibility with Devices
 - Section 8 2 Answers The Nature Of Covalent Bonding Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 8 2 Answers The Nature Of Covalent Bonding
 - Highlighting and Note-Taking Section 8 2 Answers The Nature Of Covalent Bonding
 - Interactive Elements Section 8 2 Answers The Nature Of Covalent Bonding
8. Staying Engaged with Section 8 2 Answers The Nature Of Covalent Bonding
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 8 2 Answers The Nature Of Covalent Bonding
9. Balancing eBooks and Physical Books Section 8 2 Answers The Nature Of Covalent Bonding
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 8 2 Answers The Nature Of Covalent Bonding
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section 8 2 Answers The Nature Of Covalent Bonding
 - Setting Reading Goals Section 8 2 Answers The Nature Of Covalent Bonding
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section 8 2 Answers The Nature Of Covalent Bonding
 - Fact-Checking eBook Content of Section 8 2 Answers The Nature Of Covalent Bonding
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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