

13.1**THE NATURE OF GASES****Section Review****Objectives**

- Describe the assumptions of the kinetic theory as it applies to gases
- Interpret gas pressure in terms of kinetic theory
- Define the relationship between Kelvin temperature and average kinetic energy

Vocabulary

- | | | |
|------------------|------------------------|-----------------------------|
| • kinetic energy | • vacuum | • pascal (Pa) |
| • kinetic theory | • atmospheric pressure | • standard atmosphere (atm) |
| • gas pressure | • barometer | |

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

- The kinetic theory describes the 1 of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly 2, that the particles of gas are relatively 3, move 4 of each other, and are in constant 5 motion. The 6 between particles are perfectly elastic so that the total 7 remains constant. Gas pressure results from the simultaneous collisions of billions of particles with an object. Barometers are used to measure 8 pressure. Standard conditions are defined as a temperature of 9 and a pressure of 10.
1. _____
 2. _____
 3. _____
 4. _____
 5. _____
 6. _____
 7. _____
 8. _____
 9. _____
 10. _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 11. Atmospheric pressure is 760 mm Hg.
- _____ 12. The SI unit of pressure is the pascal.

Section The Nature Of Gases Answers

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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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Chapter 10 Wave Optics 174 Chapter 11 Dual Nature of Radiation and Matter 181 Chapter 12 Atoms 188 Chapter 13 Nuclei
194 Chapter 14 Semiconductor Electronics Materials Devices and Simple Circuits 200 *Nature London* ,1873

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