

- 4 (a) (i) Balance the equation for the reaction between sodium and chlorine to produce sodium chloride by putting numbers in the spaces provided. (2)
- $$\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$$
- (ii) In an experiment to make sodium chloride, the yield is 2.5 g. The theoretical yield of sodium chloride for this experiment is 4.0 g. Calculate the percentage yield of sodium chloride in this experiment. (2)

percentage yield = _____ %

- (iii) Sodium chloride has a high melting point. Sodium chloride does not conduct electricity when solid but does conduct electricity when molten.

Complete the sentence by putting a cross (X) in the box next to your answer.

These properties show that the structure of sodium chloride is

- ☐ A ionic
- ☐ B giant molecular, covalent
- ☐ C simple molecular, covalent
- ☐ D metallic

(1)

- (iv) Calculate the relative formula mass of sodium chloride, NaCl. (relative atomic masses: Na = 23, Cl = 35.5)

(1)

relative formula mass = _____

- (b) Calculate the percentage by mass of magnesium in magnesium sulfate, MgSO_4 .

(relative atomic masses: O = 16, Mg = 24, S = 32

relative formula mass: $\text{MgSO}_4 = 120$)

(2)

percentage of magnesium = _____ %

- (c) The formula of a molecule of ethane is C_2H_6 .

- (i) Give the empirical formula of ethane.

(1)

- (ii) Ethane is a simple molecular, covalent compound.

Ethane has a low boiling point.

Explain, in terms of particles it contains, why ethane has a low boiling point.

(2)

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**Cara Flanagan, Julia Russell, Mandy
Wood, Matt Jarvis, Rob Liddle**



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