

Study Question 2a)

New Chemical Reaction

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Page:

Chemical Thermodynamic Relationships

1. Standard reaction has worked the following reactions:

a) $\text{Acetylene} + \text{Hydrogen} \rightarrow \text{Ethane}$ with $\Delta H_{\text{rxn}}^{\circ} = -210.0 \text{ kJ/mol}$ and $\Delta G_{\text{rxn}}^{\circ} = -175.7 \text{ kJ/mol}$



b) $\text{Ethane} + \text{Oxygen} \rightarrow \text{Carbon Dioxide} + \text{Water}$ with $\Delta H_{\text{rxn}}^{\circ} = -1560.7 \text{ kJ/mol}$ and $\Delta G_{\text{rxn}}^{\circ} = -1420.7 \text{ kJ/mol}$



2. Standard free energy values:

a) $\Delta G_{\text{f}}^{\circ}(\text{C}_2\text{H}_2, \text{g}) = 69.0 \text{ kJ/mol}$, $\Delta G_{\text{f}}^{\circ}(\text{H}_2, \text{g}) = 0$, $\Delta G_{\text{f}}^{\circ}(\text{C}_2\text{H}_6, \text{g}) = -67.0 \text{ kJ/mol}$

b) $\Delta G_{\text{f}}^{\circ}(\text{H}_2\text{O}, \text{l}) = -237.1 \text{ kJ/mol}$, $\Delta G_{\text{f}}^{\circ}(\text{CO}_2, \text{g}) = -394.4 \text{ kJ/mol}$

Combustion Processes

1. Standard combustion reactions for the following compounds:

a) $\text{C}_2\text{H}_2(\text{g}) + 5/2\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$ $\Delta H_{\text{rxn}}^{\circ} = -1300 \text{ kJ}$

b) $\text{C}_2\text{H}_6(\text{g}) + 7/2\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O}(\text{l})$

Equilibrium and Thermodynamic Relationships

1. Write an equation for the decomposition process:



2. Write an equation for the reaction of carbon monoxide and oxygen:



Given: $\Delta H_{\text{f}}^{\circ}(\text{CO}) = -110.5 \text{ kJ/mol}$, $\Delta H_{\text{f}}^{\circ}(\text{CO}_2) = -393.5 \text{ kJ/mol}$, $\Delta H_{\text{f}}^{\circ}(\text{H}_2\text{O}) = -285.8 \text{ kJ/mol}$

Skill Practice 36 Answers

Murray Shukyn



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