

Half Life and Radioactive Decay

Name: _____

Instructions: The *half-life* of a radioactive substance is a characteristic constant. It measures the time it takes for a given amount of the substance to become reduced by half as a consequence of decay, and therefore, the emission of radiation. Scientists use half-life to calculate amount of radioactive isotopes present in a sample.

Sample Half-life Problem: A 100 gram sample of carbon-14 undergoes radioactive decay until 12.5 grams of carbon-14 remains in the original sample. This process took 17,190 years. How many half-lives have passed? What is the half-life of ^{14}C ?

Step 1: Make a data table with the original isotope, half-lives and time passed as its column headers.

Isotope (g)	Number of Half-lives Passed	Time Passed (Years)
100g	0	0
50	1	
25	2	

Step 2: Using the half-life principle, deduct the correct amount of isotope and record the number of half-lives passed until you reach the given number of years.

Amount of Isotope (g)	Number of Half-lives Passed	Time Passed (Years)
100g	0	

Step 3: Divide the number of year passed by the number of half-lives passed: $17,190/5 = 3438$ years in one half-life. Fill in the remainder of the chart.

Amount of Isotope (g)	Number of Half-lives Passed	Time Passed (Years)
100g	0	0
50	1	3438
25	2	6876
12.5	3	10314

Half-life Problems:

- What is the half-life of ^{235}U if a 500g sample decays to 30.25 grams in 63 hours?
- How much of a 100 g sample of ^{14}C would remain after 37 years if the half-life is 12,35 years?
- The half-life of ^{239}Pu is 1.4×10^4 years. If there are 6.25 g of the sample left after 2.8×10^4 years, how many grams were in the original sample?
- The half-life of Iodine-131 is 8.04 days. How long will it take a 200 g sample to decay into 6.25g?

Radioactive Decay And Half Life Worksheet Answers

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web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as 6.02×10^{23} particles this is the avogadro number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a mole is always 6.02×10^{23} but as different particles have different mass their molar mass varies

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web the molar volume has the si unit of cubic metres per mole $\text{m}^3 \text{mol}^{-1}$ although it is more typical to use the units cubic decimetres per mole $\text{dm}^3 \text{mol}^{-1}$ for gases and cubic centimetres per mole $\text{cm}^3 \text{mol}^{-1}$ for liquids and solids

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web zero degrees celsius or 273 k is defined as standard temperature and one atmosphere or 760 mm hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure stp is 22.4 l

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web at standard temperature and pressure stp the molar volume V_m is the volume occupied by one mole of a chemical element or a chemical compound it can be calculated by dividing the molar mass M by mass density ρ molar gas volume is one mole of any gas at a specific temperature and pressure has a fixed volume molar volume formula

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web this volume is known as the molar gas volume V_m the volume of a gas may be calculated from its number of moles using volume of gas moles V_m rearranging this moles of gas frac

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web jun 8 2023 $V = 1.00 \text{ mol} \times 0.08206 \text{ latm mol}^{-1} \text{K}^{-1} \times 273 \text{K} = 22.4 \text{ l}$ in other words 1 mole of a gas will occupy 22.4 l at stp assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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mole of any gas at room temperature and pressure occupies 24 dm^3 or $24\,000 \text{ cm}^3$ this value is known as the molar volume
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