

5. Predict which system at equilibrium will contain the maximum amount of product and which system will contain the maximum amount of reactant?



Ans: The magnitude of K_c can give us some information about the reactant and product concentrations at equilibrium:

- If K_c is very large, ~ 1000 or more, we will have mostly product species present at equilibrium.
- If K_c is very small, ~ 0.001 or less, we will have mostly reactant species present at equilibrium.
- If K_c is between 0.001 and 1000 , we will have a significant concentration of both reactant and product species present at equilibrium.

By using these guidelines, we can quickly estimate whether a reaction will strongly favour the forward direction to make products very large K_c – strongly favour the backward direction to make reactants very small K_c or somewhere in between.

Therefore, system (b) will contain the maximum amount of product.

SECTION 'D' : NUMERICAL

1. Dinitrogen tetraoxide N_2O_4 decomposed into nitrogen dioxide NO_2 in a reversible reaction. Derive equilibrium constant expression for the reaction of decomposition. Also, interpret the unit of K_c for balanced chemical reversible reaction.

Ans: Equation:-



Equilibrium Constant Expression: $K_c = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$

Unit of K_c : $K_c = \frac{[\text{mol}.\text{dm}^{-3}]^2}{[\text{mol}.\text{dm}^{-3}]}$
 $= \text{mol}.\text{dm}^{-3}$

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Clemens Wendtner



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