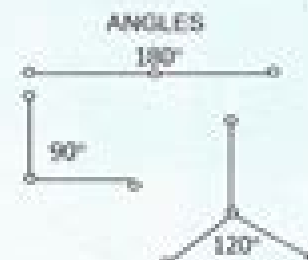


CHEMISTRY - MOLECULAR GEOMETRY ACTIVITY

1. TWO BONDED PAIRS		Bond Angle: 180° Geometry: <u>linear</u> Type: <u>linear</u>	1 <u>= lone pair</u> UNSHARED PAIR OF ELECTRONS						
2. THREE BONDED PAIRS		Bond Angle: 120° Geometry: <u>Trigonal Planar</u> Type: <u>Planar</u>	2 UNSHARED PAIR OF ELECTRONS		Bond Angle: $< 120^\circ$ Geometry: <u>Planar</u> Type: <u>Bent</u>	<u>polar</u>			
3. FOUR BONDED PAIRS		Bond Angle: 109.5° Geometry: <u>Tetrahedral</u> Type: <u>Tetrahedral</u>	3 UNSHARED PAIR OF ELECTRONS		Bond Angle: $< 109.5^\circ$ Geometry: <u>Trigonal Bipyramidal</u> Type: <u>Bent</u>	<u>polar</u>			
4. FIVE BONDED PAIRS		Bond Angle: $120/90^\circ$ Geometry: <u>Bipyramidal</u> Type: <u>Bipyramidal</u>	4 UNSHARED PAIR OF ELECTRONS		Bond Angle: $< 120^\circ$ Geometry: <u>Trigonal Bipyramidal</u> Type: <u>Bent</u>	<u>polar</u>			
5. SIX BONDED PAIRS		Bond Angle: 90° Geometry: <u>octahedral</u> Type: <u>octahedral</u>	5 UNSHARED PAIR OF ELECTRONS		Bond Angle: 90° Geometry: <u>Square Planar</u> Type: <u>Bent</u>	<u>polar</u>			

TYPE of Structure
Description:
M=Central Atom
X=Substituent Atom
E=Lone Pair Electrons



- Draw a diagram of each of the geometries possible in the box provided.
- List the Bond Angles and Geometries and Type of Structure
- How many atoms are bonded to the central atom in a molecule that has:
a. a linear shape 2 b. a trigonal planar shape 3 c. a tetrahedral shape 4 d. a trigonal bipyramidal shape 5 e. an octahedral shape 6 ?
- What other geometries are possible due to lone pairs ?
- What is the maximum bond angle for any geometry? 180° What is the minimum bond angle for any geometry? 90° ΔEN ;
What influences restrict the maximum and minimum bond angles? * lone pair e⁻, * bonding pairs, ΔEN ;
- Under what conditions is it possible to have the same geometry but different number of electrons? bent, trigonal bipyramidal
Give two examples: 2-bonded/1 lone = bent ; 2-bonded/1 lone = bent ; 4-bonded/1 lone = bent ; 2-bonded/1 lone = bent.

Vsepr Theory Yahoo Answers

Zhe Xu



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